

APPENDIX C

ELECTROMOTIVE FORCE SERIES

Table C.1*
Half Cell Reactions and Standard Potentials
(Engineering Sign Convention)

Half Cell Anode Reactions†	E^0 (Potential in Volts)†
Au \rightarrow Au ³⁺ + 3e ⁻	+1.50
Ag \rightarrow Ag ¹⁺ + 1e ⁻	+0.80
Cu \rightarrow Cu ²⁺ + 2e ⁻	+0.34
Sn \rightarrow Sn ²⁺ + 2e ⁻	-0.14
Fe \rightarrow Fe ²⁺ + 2e ⁻	-0.44
Zn \rightarrow Zn ²⁺ + 2e ⁻	-0.76
Al \rightarrow Al ³⁺ + 3e ⁻	-1.66
Li \rightarrow Li ¹⁺ + 1e ⁻	-3.05

* Askeland, D.R., 1989, Table 20-1, p. 784.

† Potential relative to hydrogen electrode. For cathode reaction, reverse the anode reaction and change the sign on E^0 .