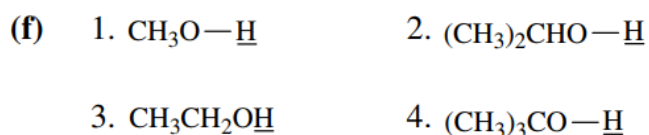
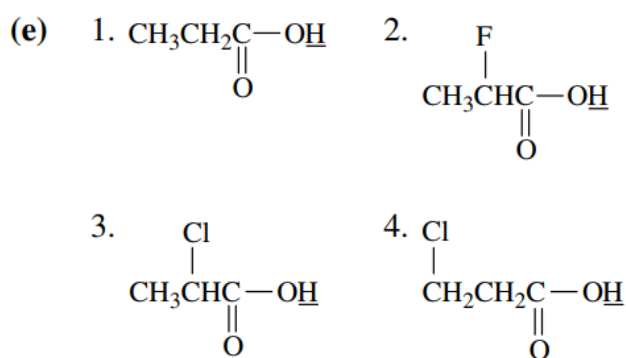
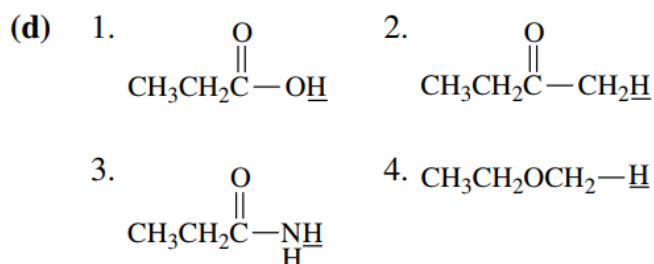
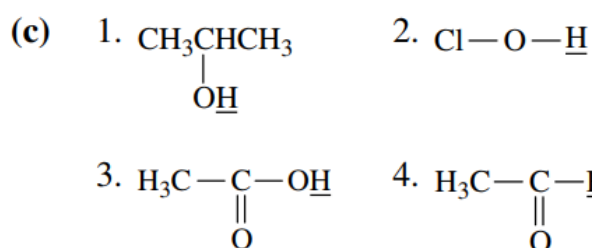
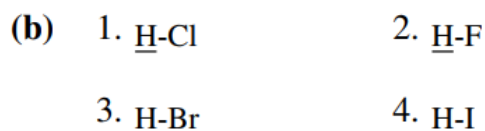
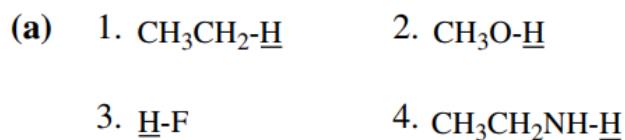
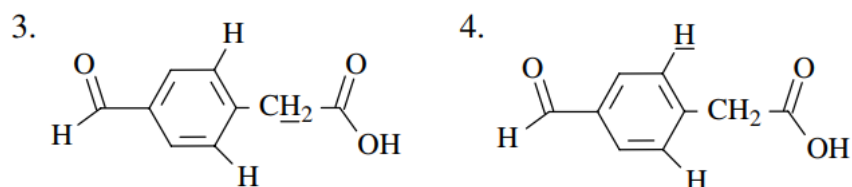
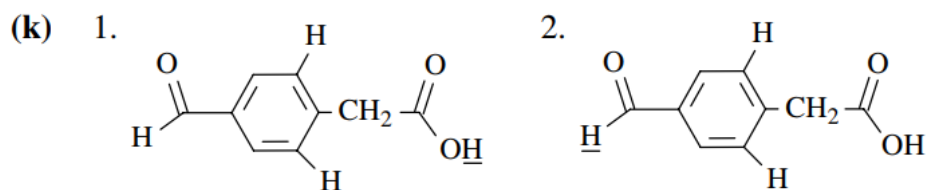
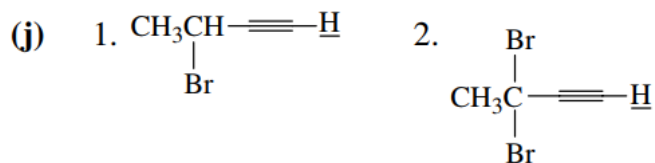
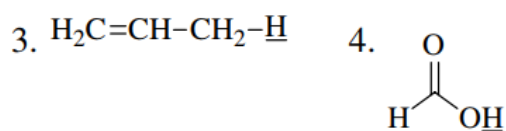
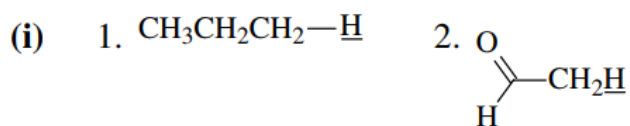
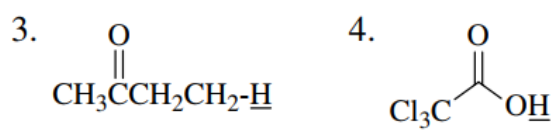
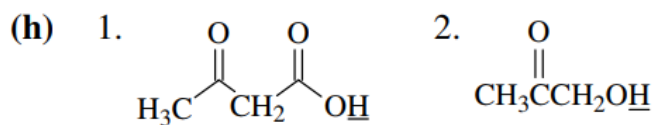
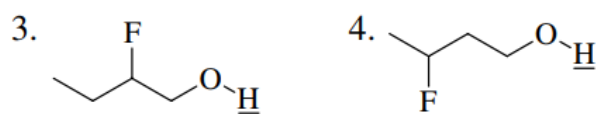
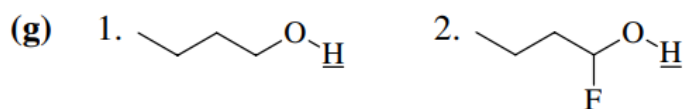


Lista de exercícios Química Orgânica 2 Farmácia Noturno 2024: Acidez e Basicidade

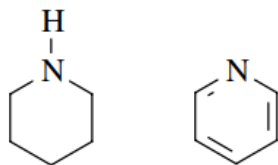
1- Coloque em ordem crescente de acidez os átomos de hidrogênio destacados a seguir:



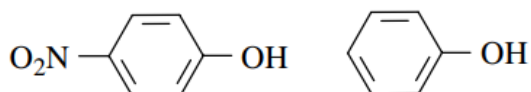


2- Explique as seguintes afirmações abaixo:

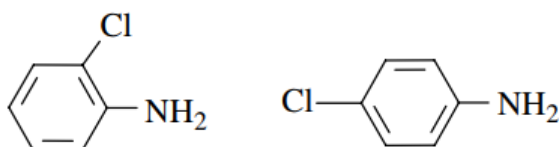
- a) Amidas tem a tendência de serem mais protonadas no oxigênio do que no nitrogênio
- b) Éteres são bases de Lewis melhores que cetonas
- c) BF_3 é um ácido de Lewis mais forte que $(\text{CH}_3\text{O})_3\text{B}$
- d) A piperidina é uma base mais forte que a piridina.



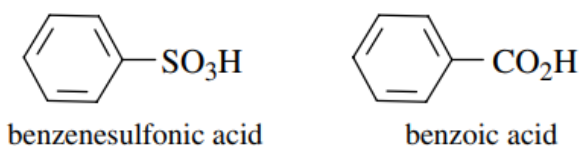
- e) p-Nitrofenol tem $\text{pK}_a \approx 8$ enquanto tem $\text{pK}_a \approx 10$



- f) o-cloroanilina é uma base mais fraca que p-cloroanilina



- g) Ácido benzenosulfônico é um ácido mais forte que o ácido benzoico



3- Dados experimentais apontam que entre ésteres derivados de ácidos benzoicos a constante de ionização (k_{rate}) é inversamente proporcional ao pK_a do respectivo ácido. A reação ocorre como apresentado no esquema abaixo. Explique este fenômeno.

