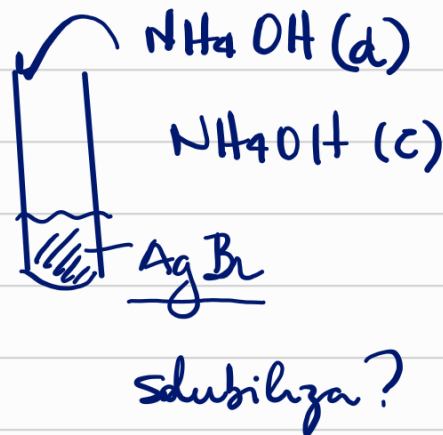
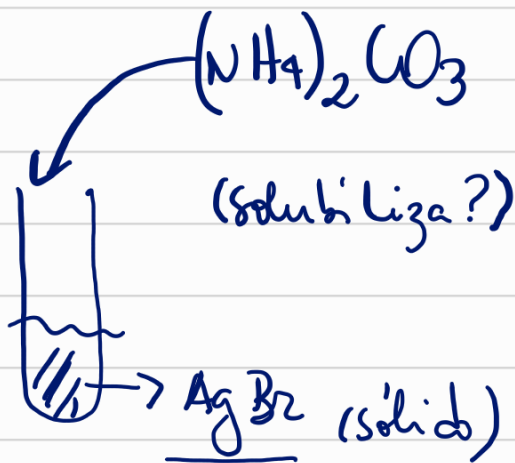
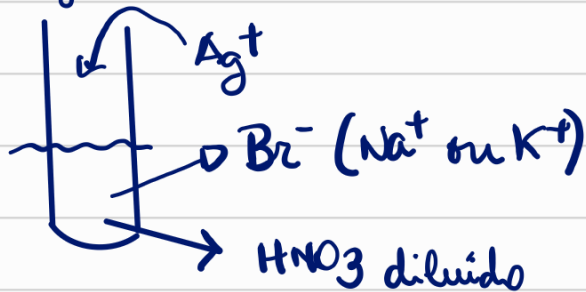
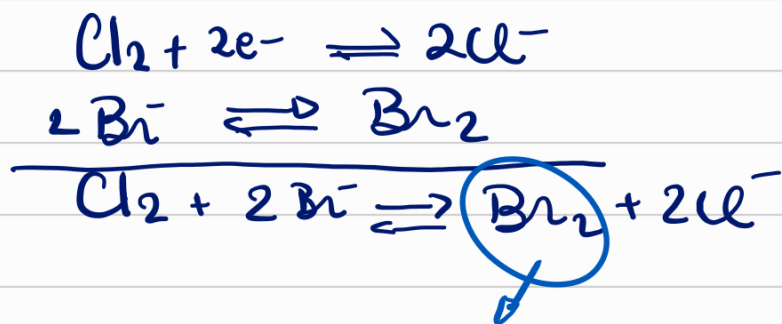
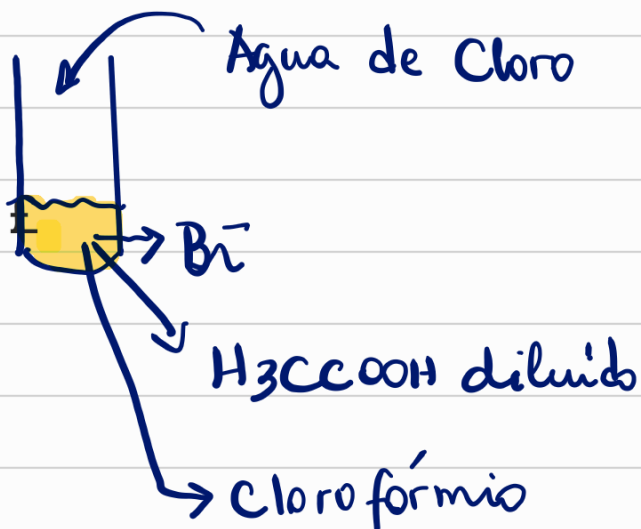


Reações dos íons Br^-

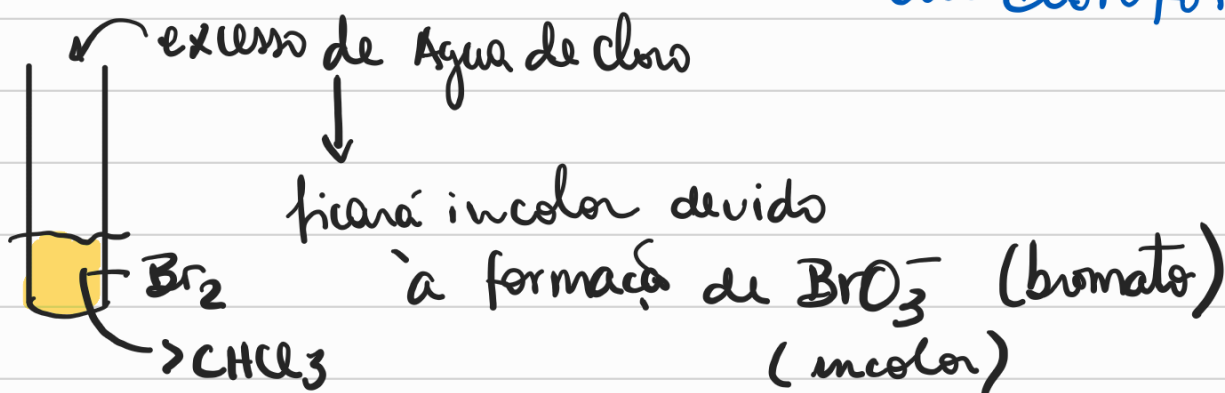
1. AgNO_3



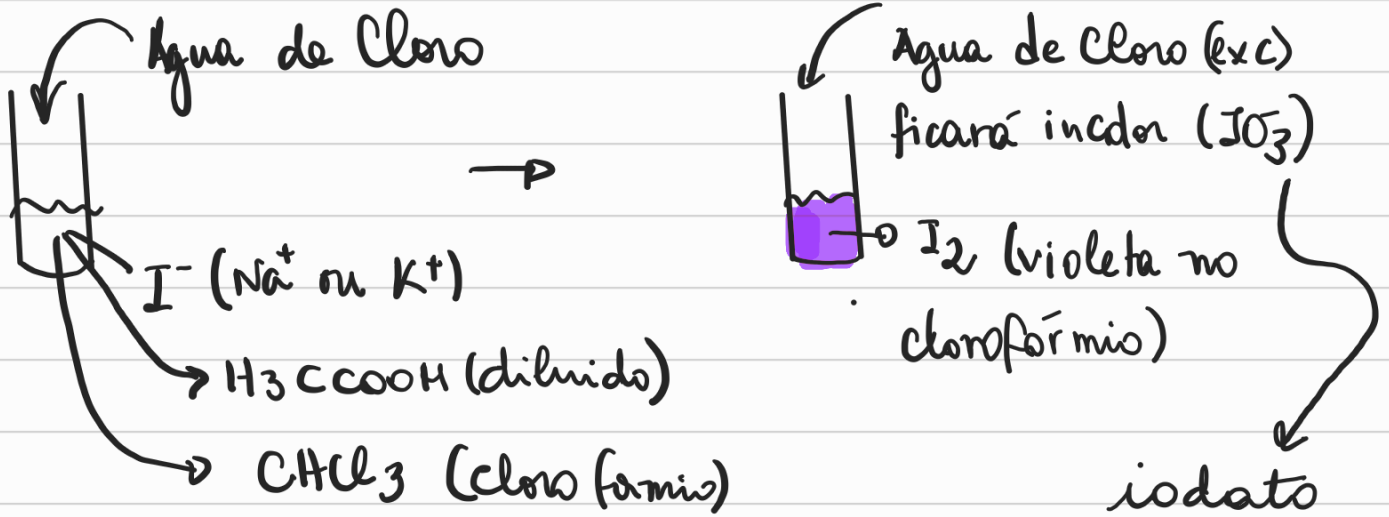
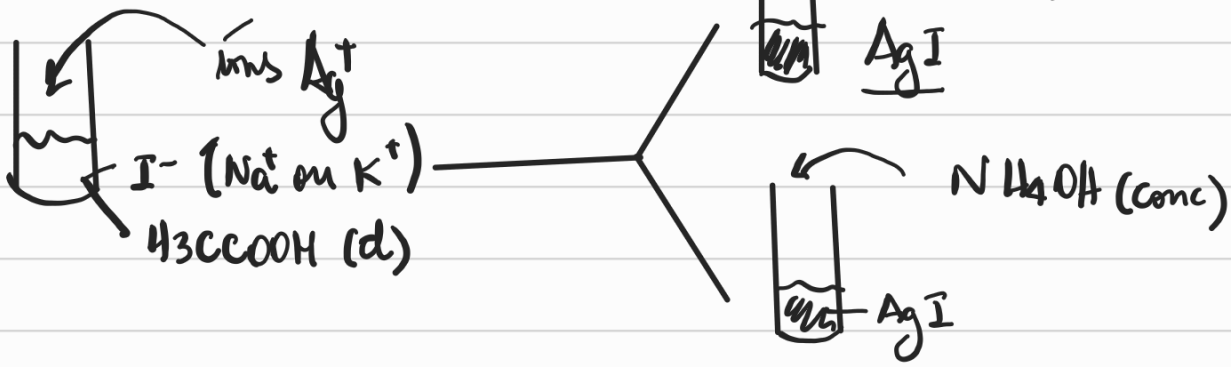
2. Água de Cloro em meio ácido (H_3CCOOH diluido)



alaranjado quando extraído em cloroformio

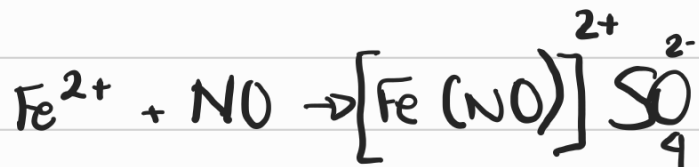
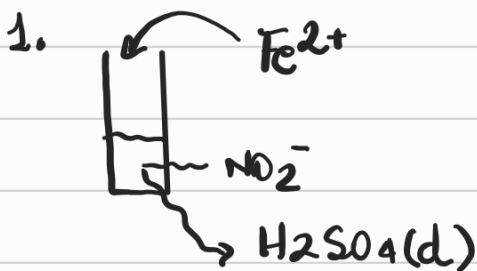


Reações dos Íons I⁻

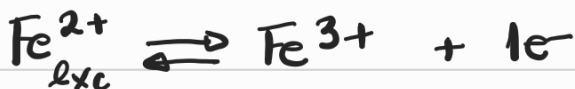


Reações dos Íons NO₂⁻

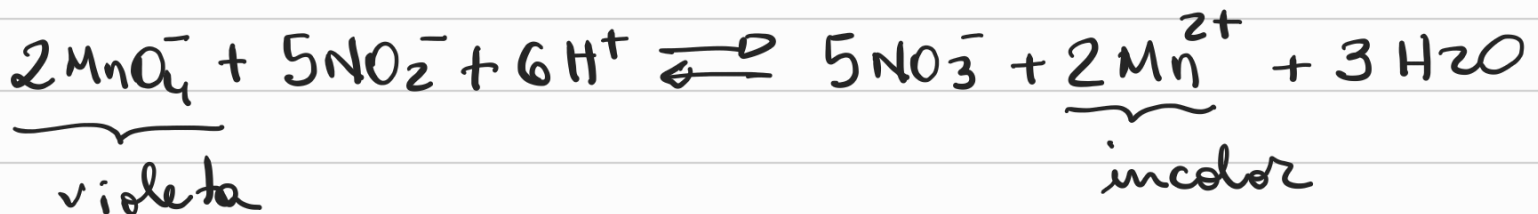
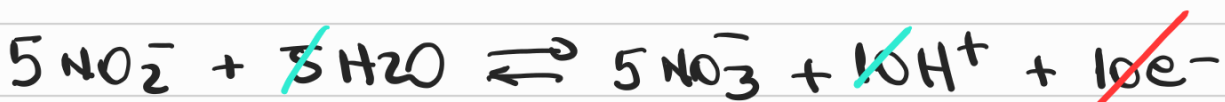
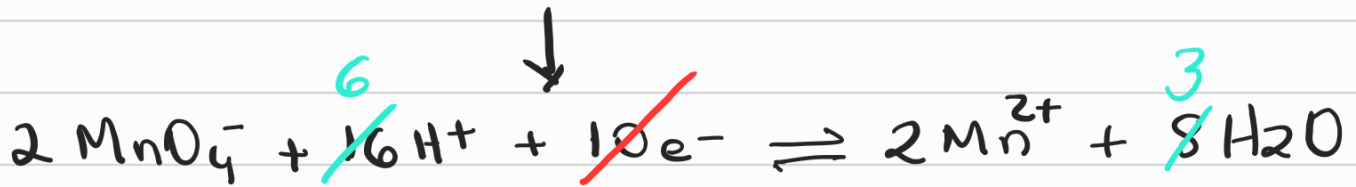
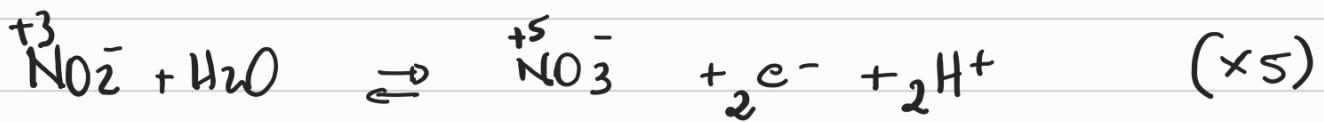
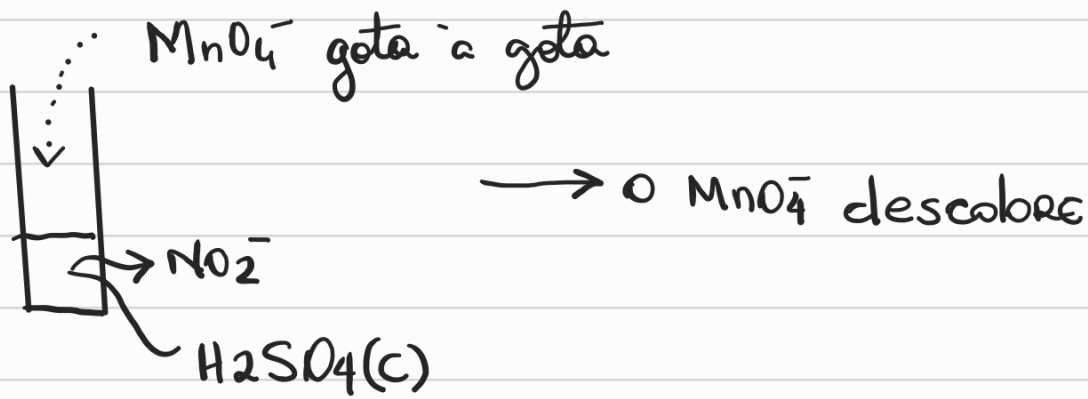
→ usar sais de Na⁺ ou K⁺



Sulfato de ferro
nitrosilo

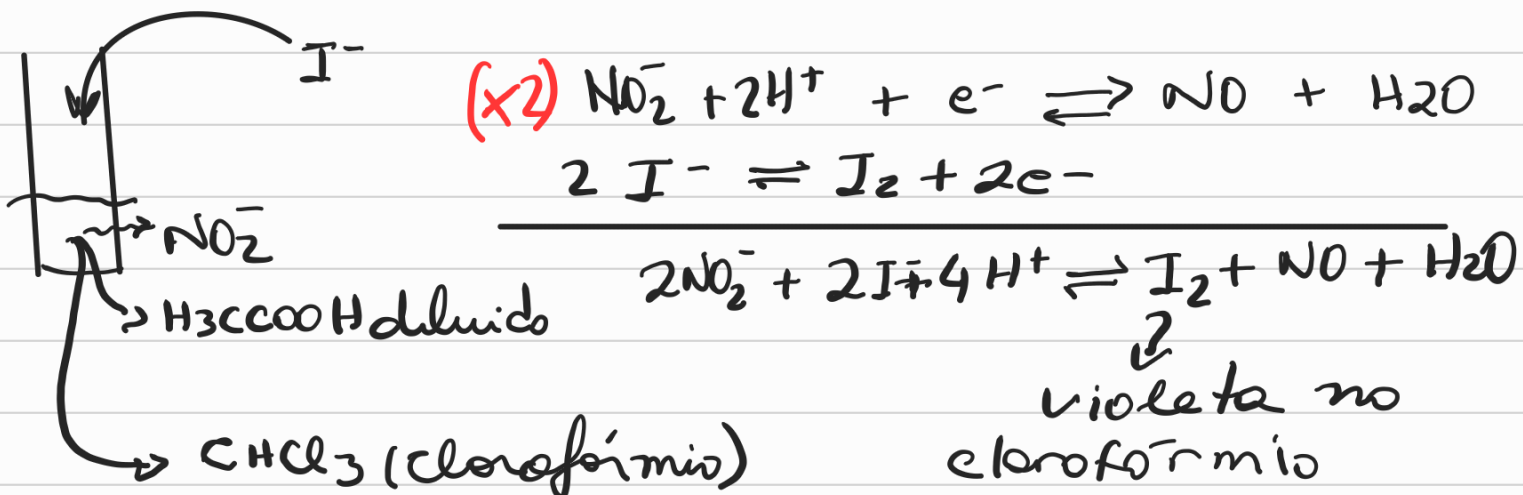


2. NO_2^- na presença de $\text{H}_2\text{SO}_4(\text{c})$ e KMnO_4

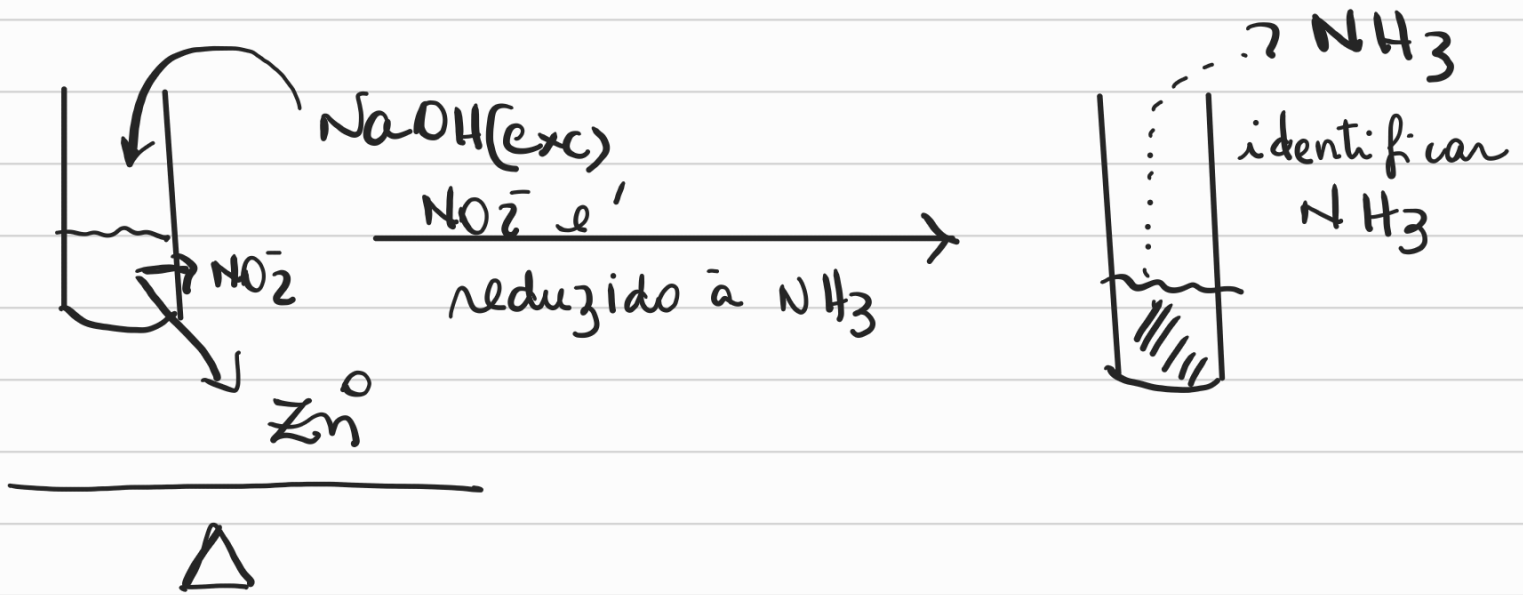


3. NO_2^- em meio ácido na presença de I^-

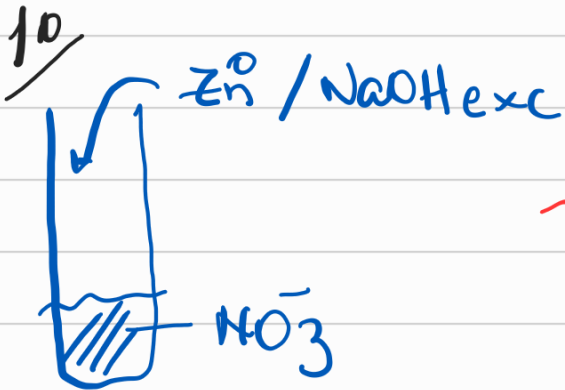
(acetato de sódio ou potássio)



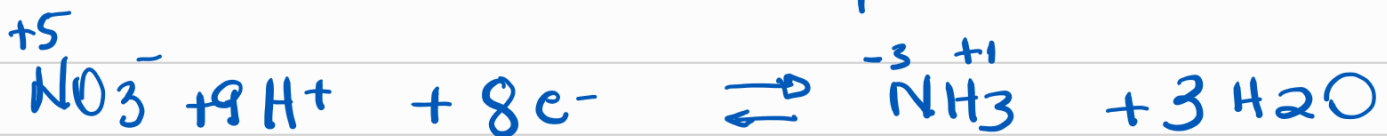
4. NO_2^- em meio alcalino na presença de Zn^0

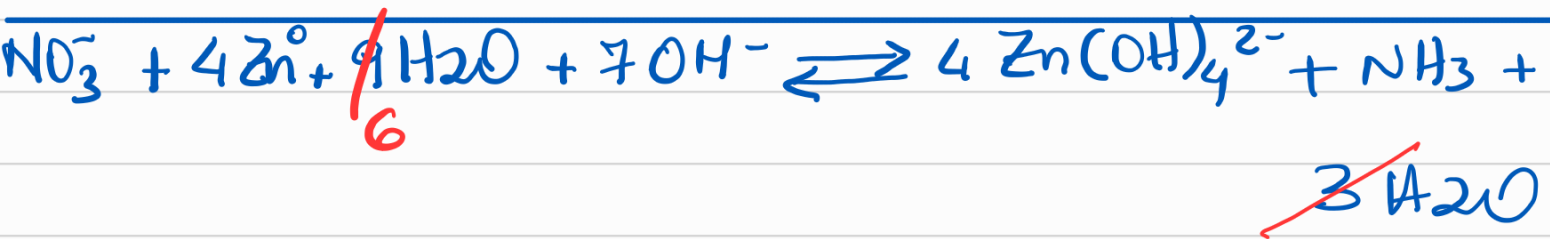
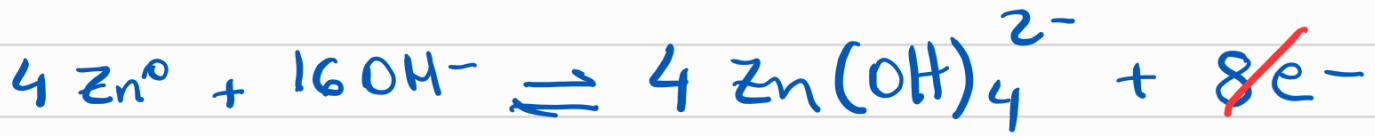


Reações do íon NO_3^-



→ identificar NH_3 na saída do tubo de ensaio.





A identificação de NH_3 comprova a existência de NO_3^- no tubo. **Cuidado**

NO_2^- apresenta reação exatamente igual. Esta prova só pode ser utilizada na ausência de NO_2^- .

Eliminação de NO_2^- é feita com adição de uréia, ácido clorídrico diluído e aquecimento.

2º Prova do Anel: Sulfato ferroso amoniacal ($\text{Fe}(\text{NH}_4)_2(\text{SO}_4)_2$) em meio de ácido sulfúrico concentrado

a) Adicionar H_2O em um tubo de ensaio. Acidificar com ácido sulfúrico diluído e dissolver totalmente, uma ponta de espátula de sulfato ferroso amoniacal. Em seguida

adicionar o sal de nitrato e homogeneizar.
Vá até a cabeça, incline o tubo de ensaio
& deixe escorrer lentamente, pelas paredes do
tubo, Ácido sulfúrico concentrado.

