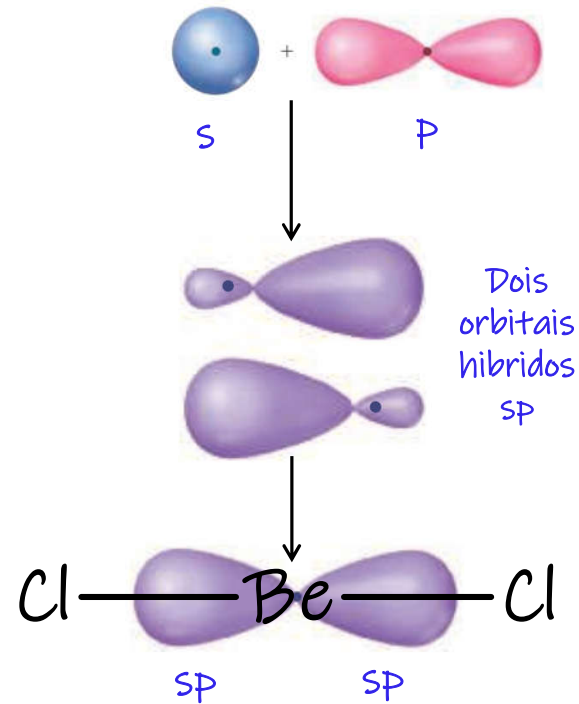
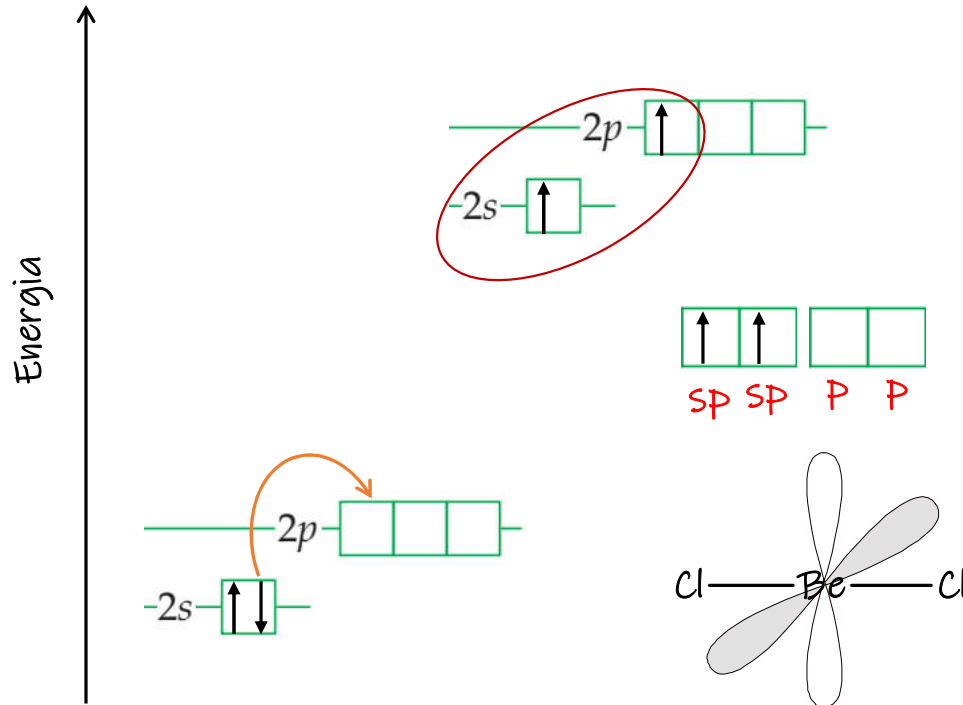
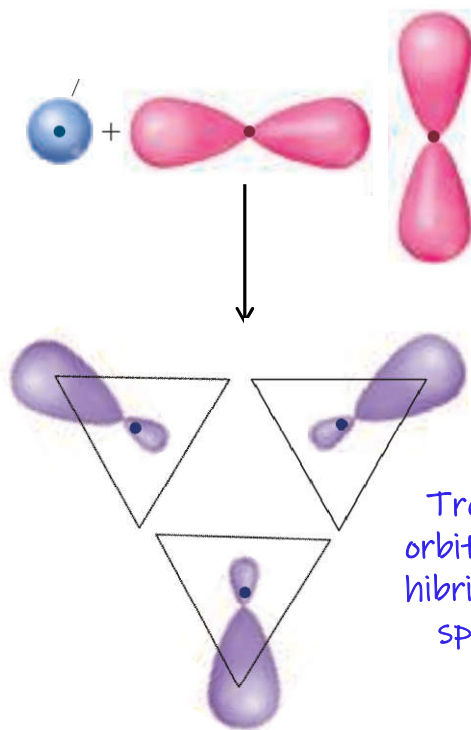
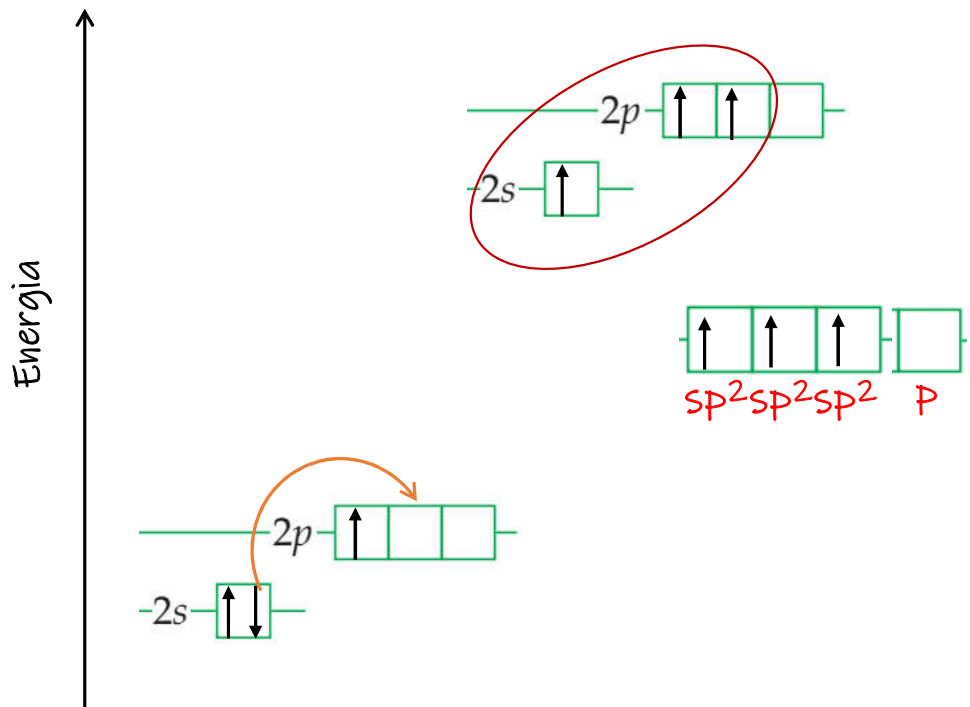
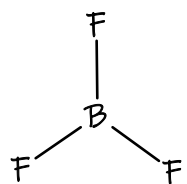
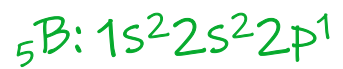


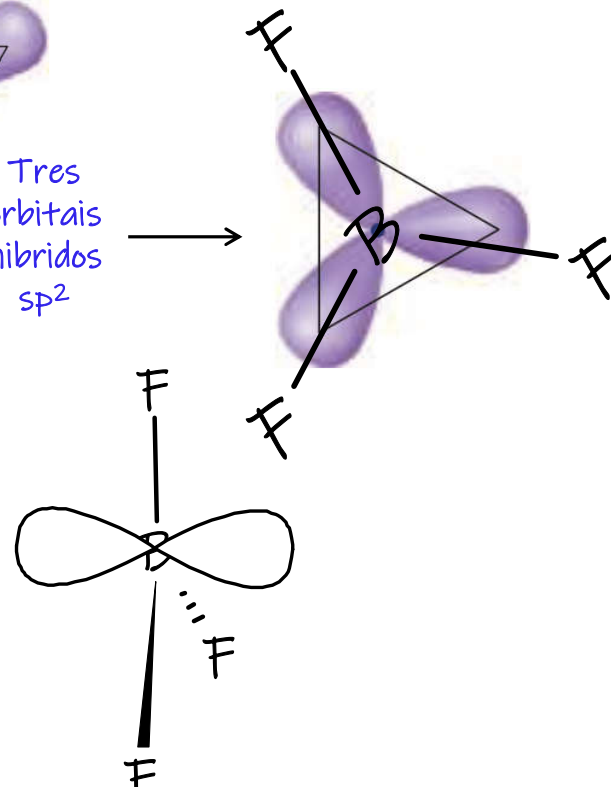
Aula 6 QE

Orbitais e Hibridização

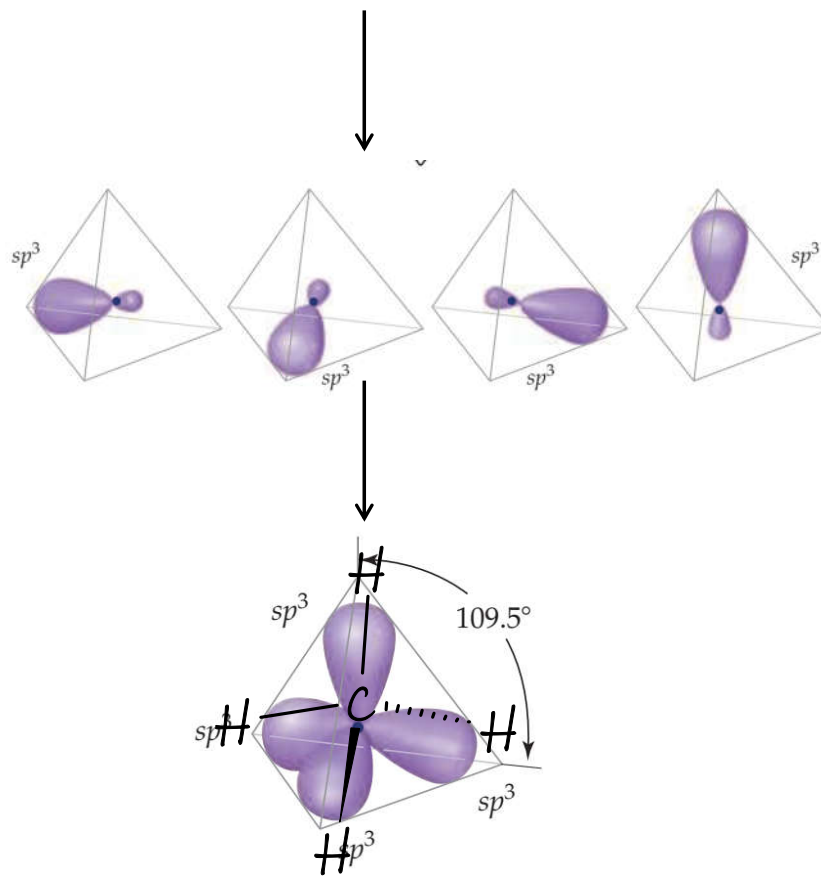
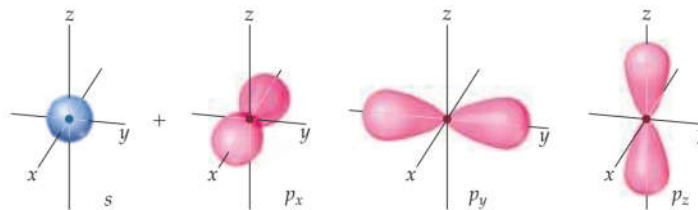
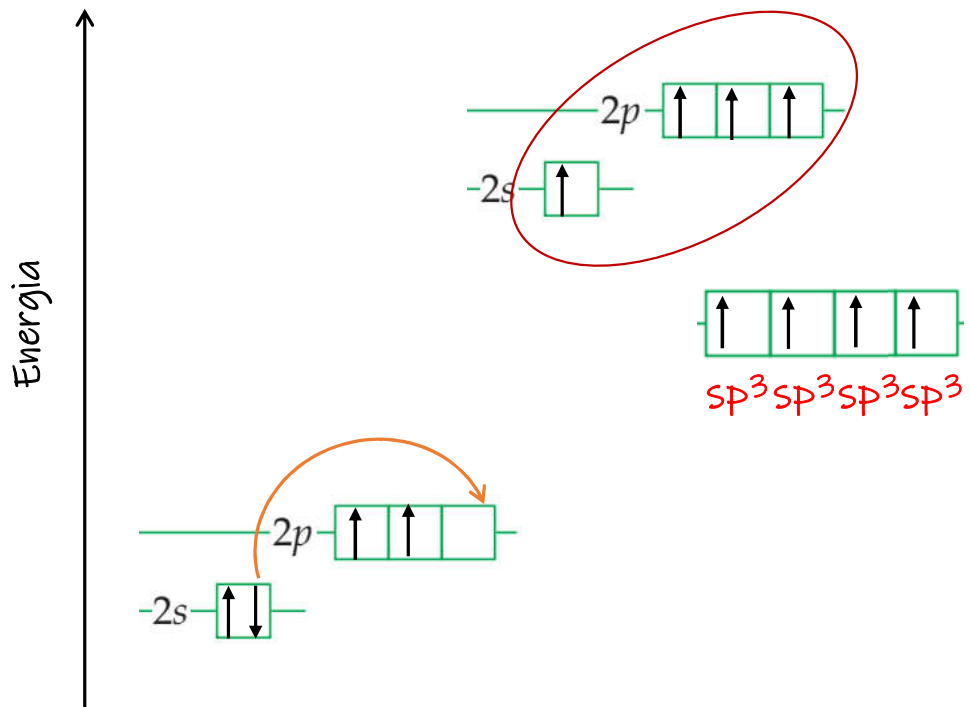
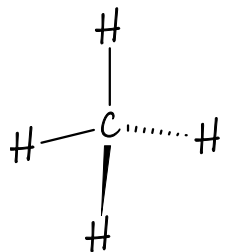




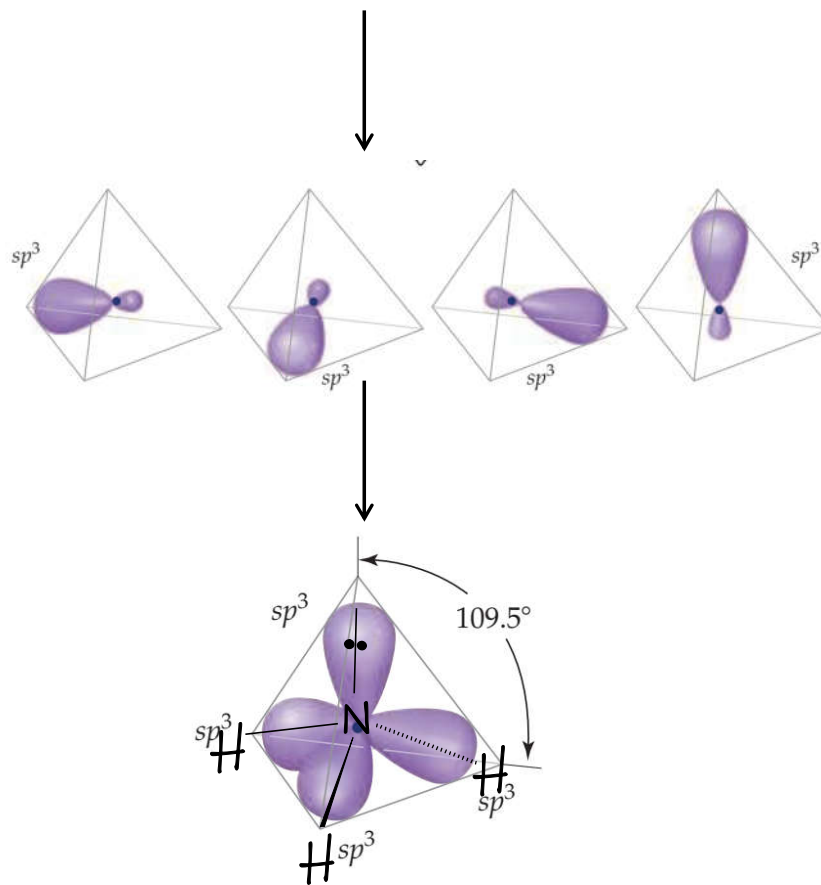
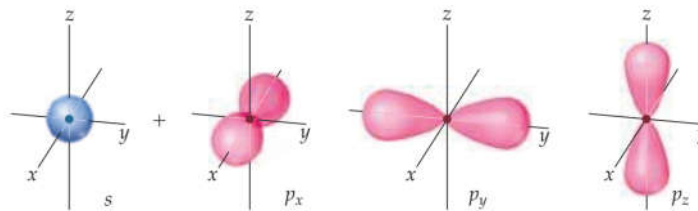
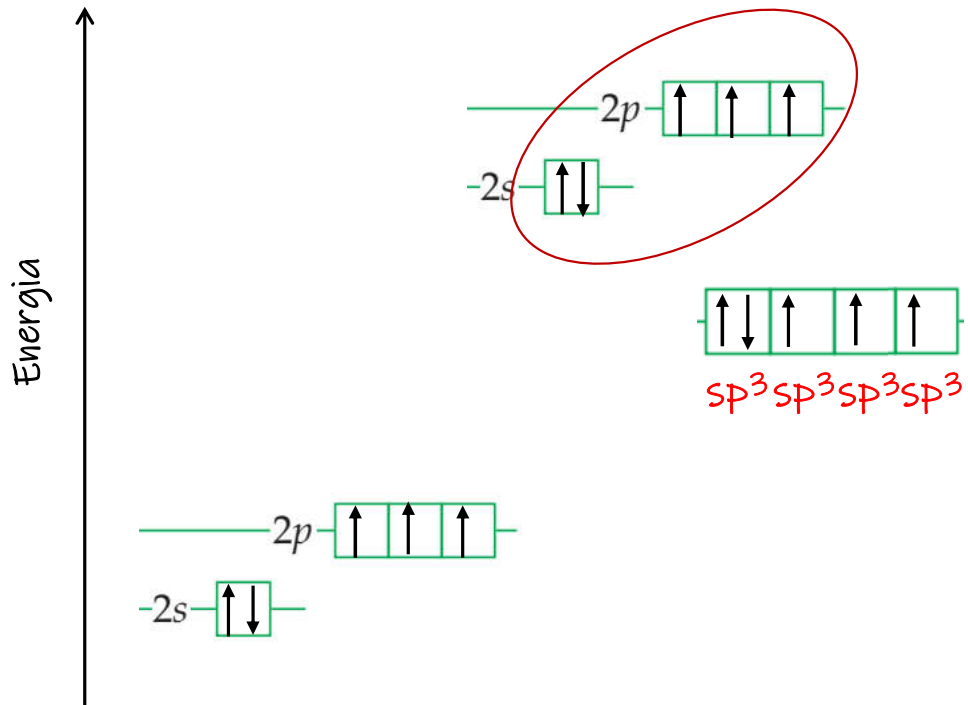
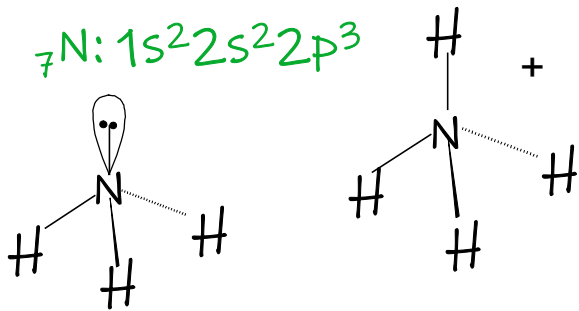
Tres orbitais hibridos sp^2



${}^6\text{C}: 1s^2 2s^2 2p^2$

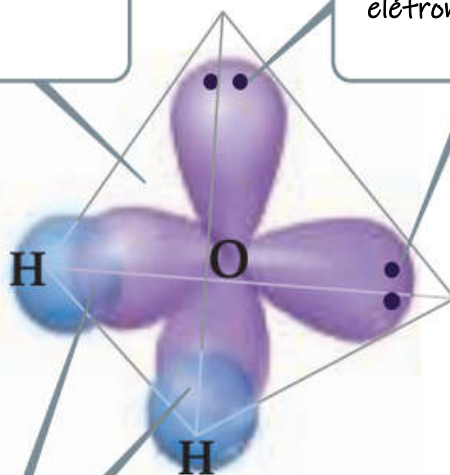


${}^7\text{N}: 1s^2 2s^2 2p^3$

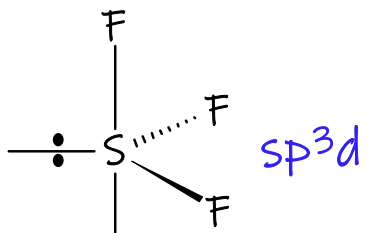


arranjo tetraédrico para 4 orbitais sp^3 do oxigênio

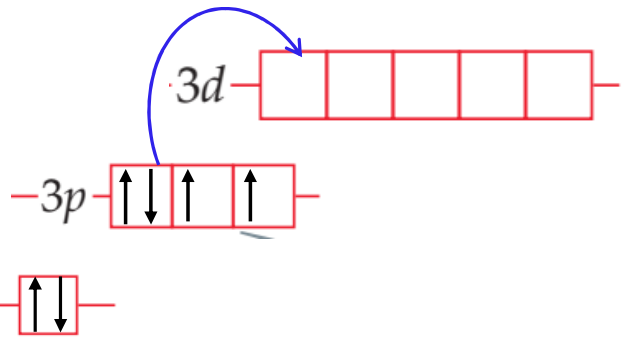
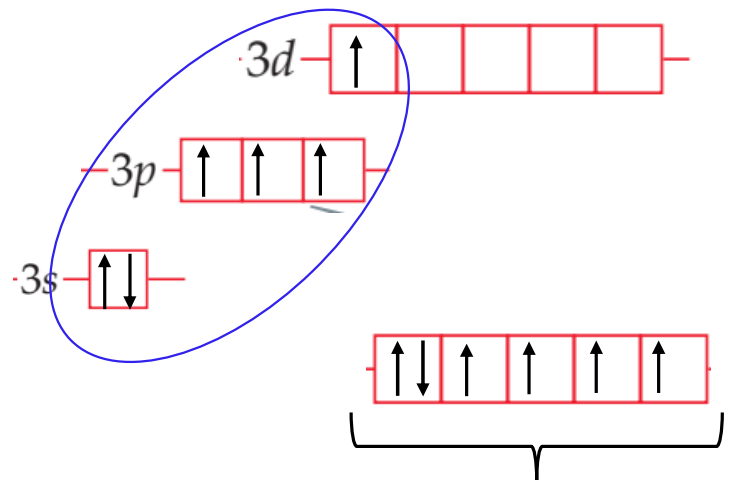
2 orbitais sp^3 contêm pares de elétrons isolados



ligações simples O-H

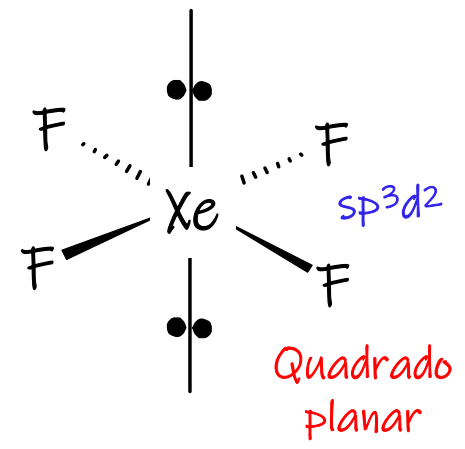
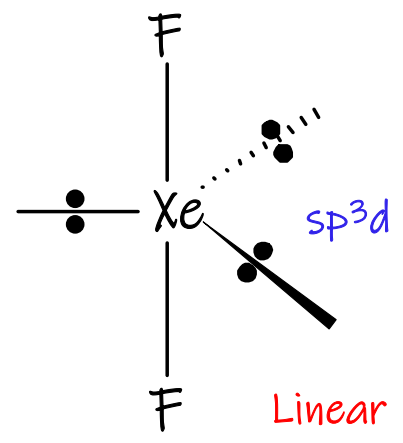


Energia

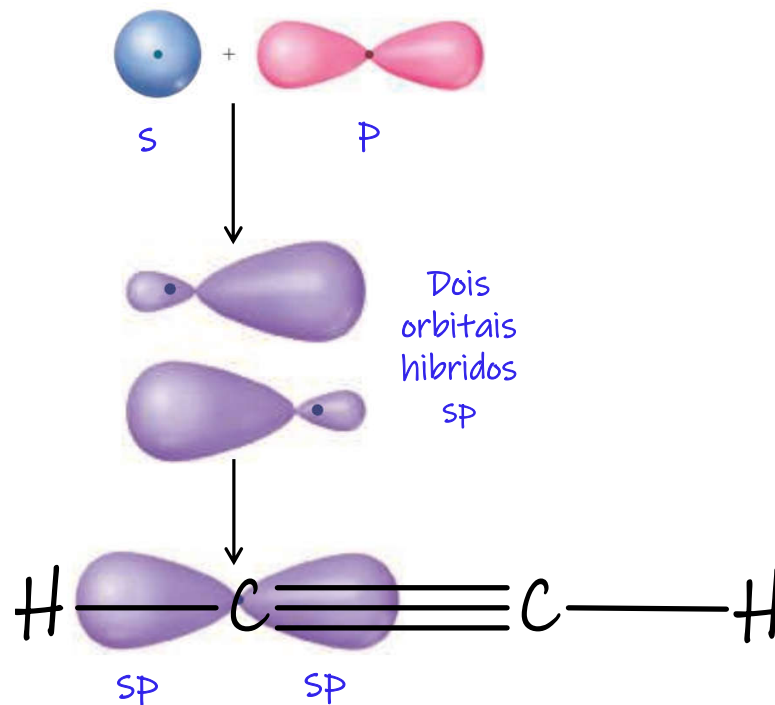
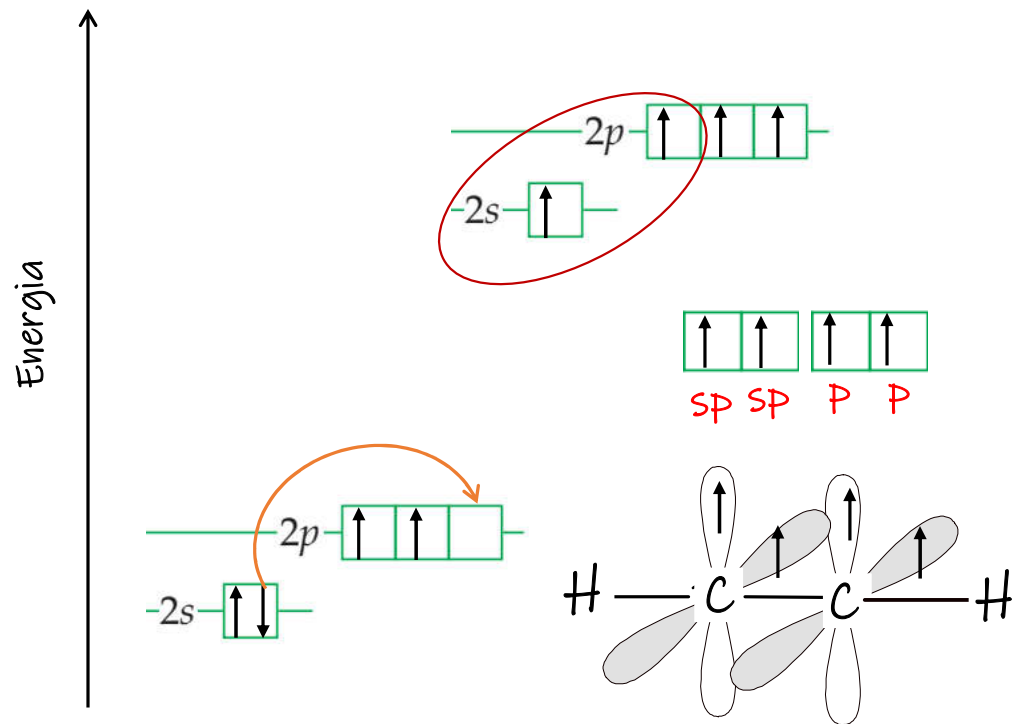
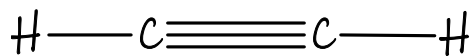
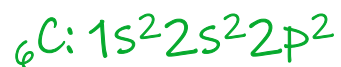


5 orbitais híbridos sp^3d

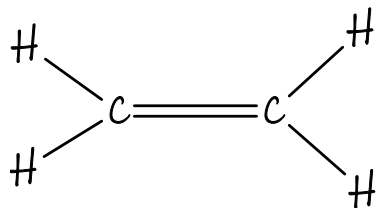
$54Xe: [Kr]5s^24d^{10}5p^6$



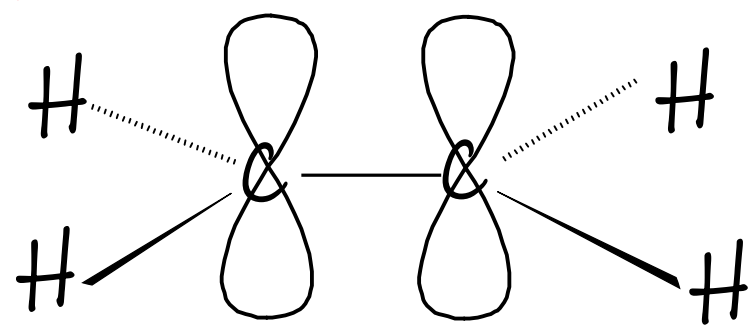
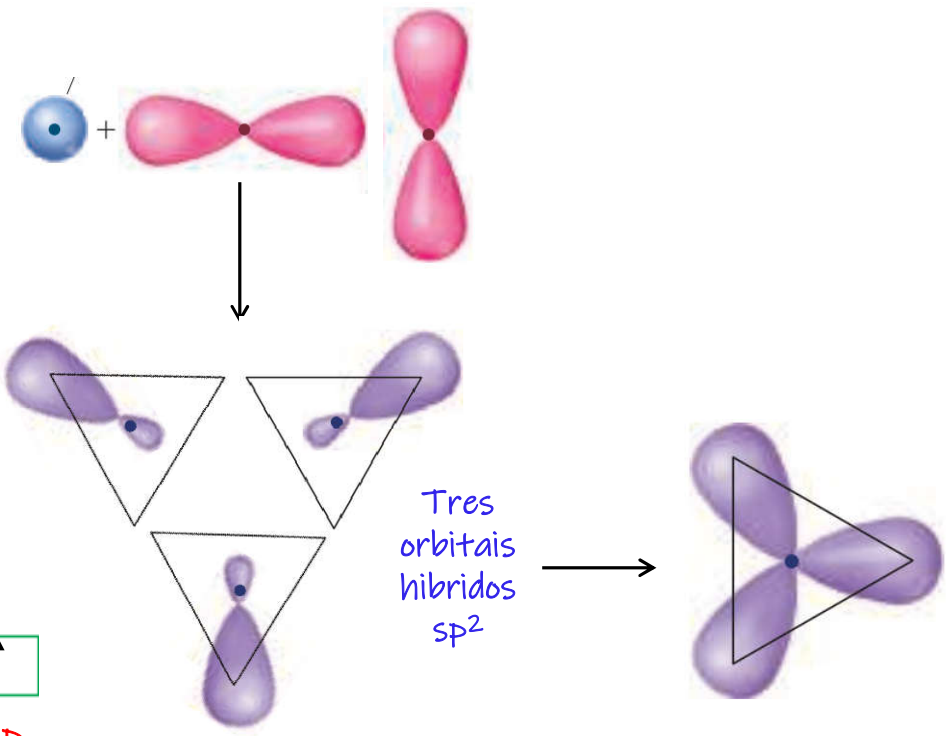
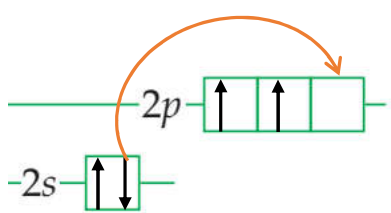
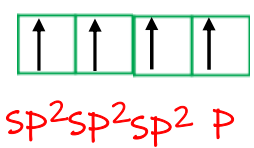
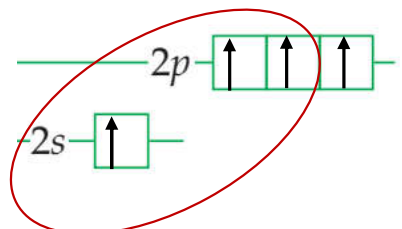
Ligações Múltiplas



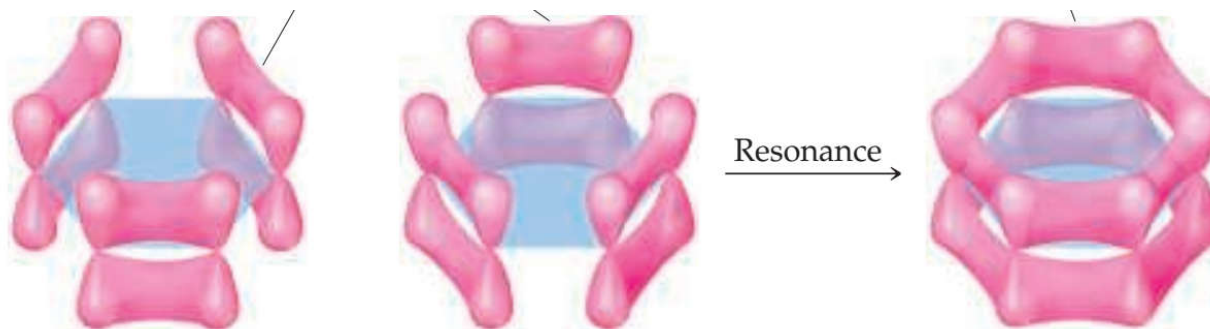
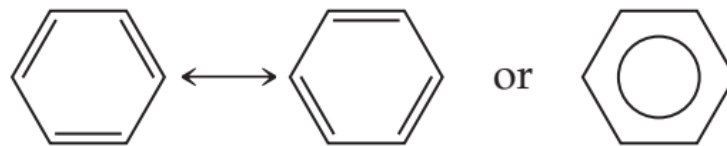
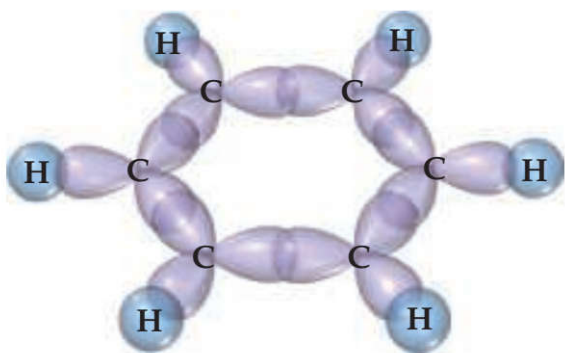
${}^6\text{C}: 1s^2 2s^2 2p^2$

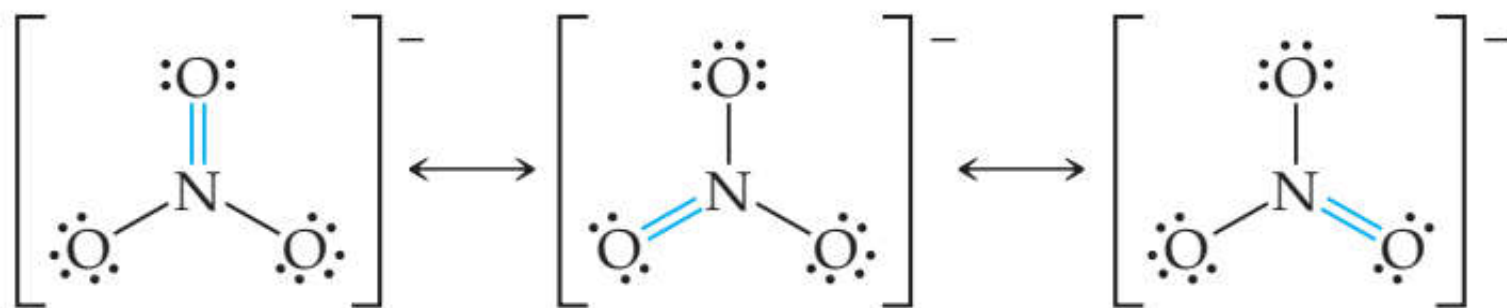


Energia ↑



Ressonância, deslocalização π

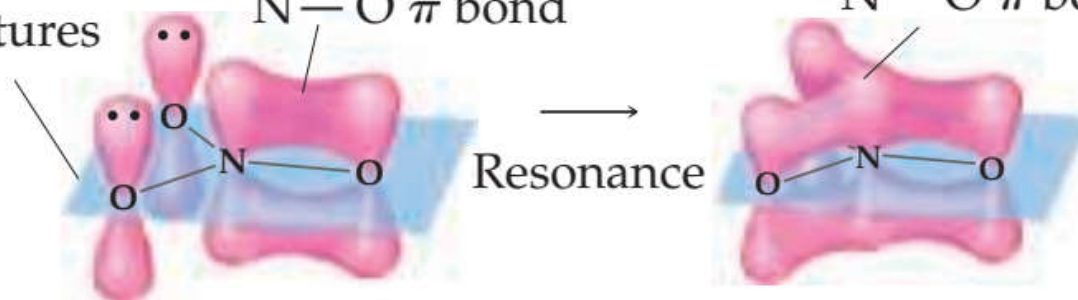




One of three
resonance
structures

Localized
N—O π bond

Delocalized
N—O π bonds



Ressonância, deslocalização π

