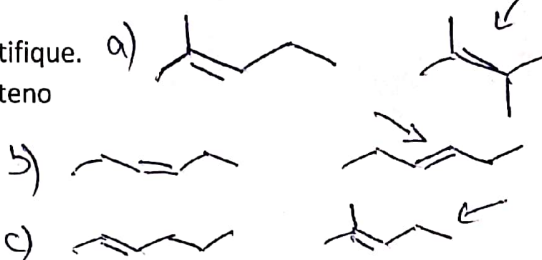


1) Qual alceno mais estável em cada par? Justifique.

a) 2-metil-2-penteno ou 2,3-dimetil-2-buteno

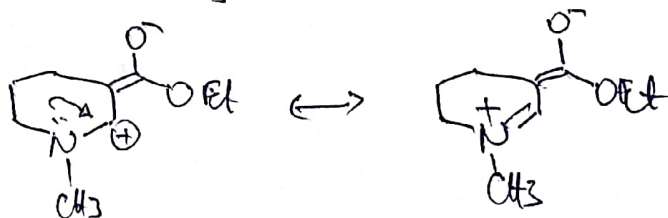
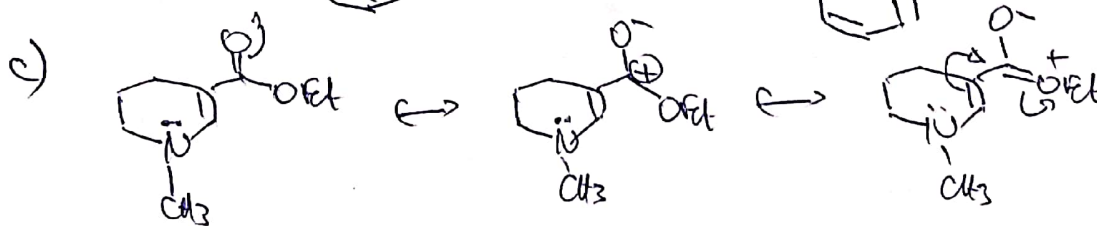
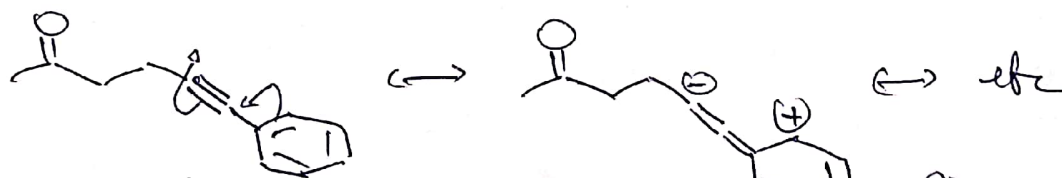
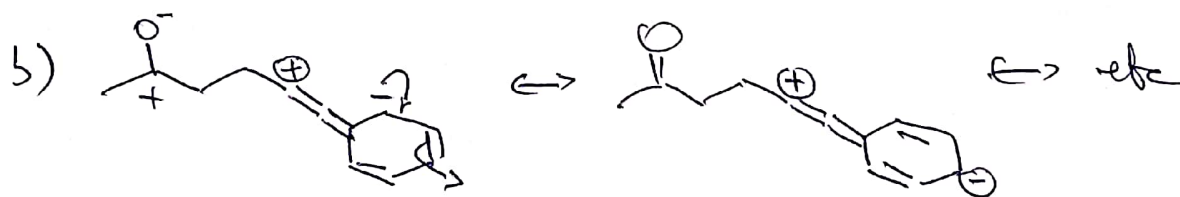
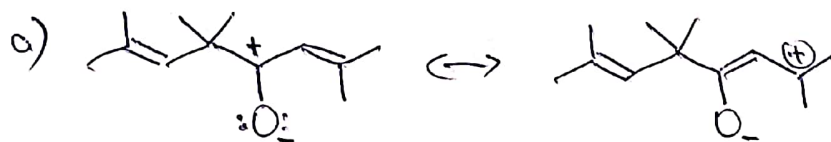
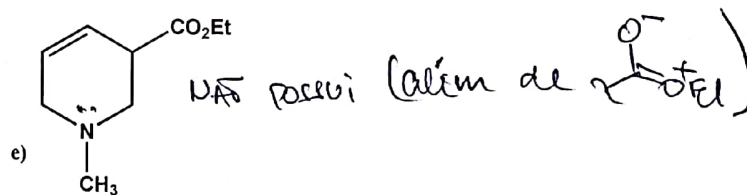
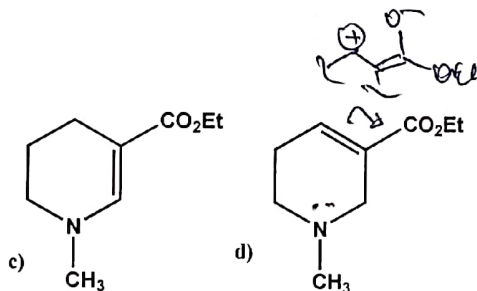
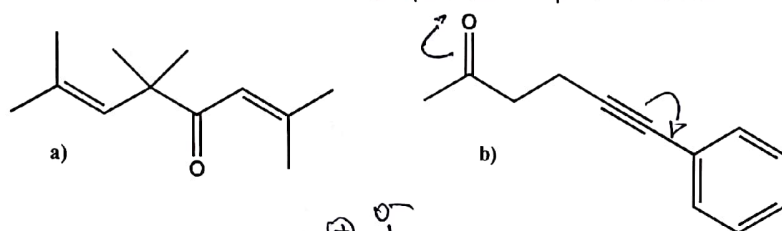
b) cis-3-hexeno ou trans-3-hexeno

c) trans-2-hexeno ou 2-metil-2-penteno.

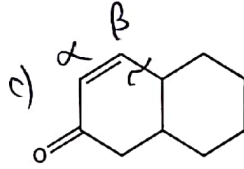
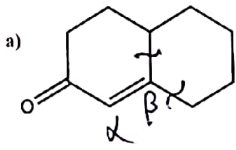


As olefinas mais substituídas ou as trans são em geral mais estabilizadas. As causas estão associadas ao mais efetivos efeitos de hiperconjugação ou ao menos efeito estérico quando os substituintes estão em trans.

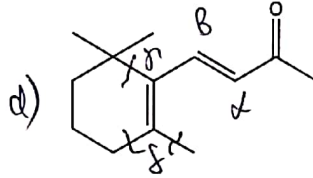
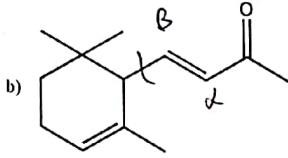
2) Represente estruturas de ressonância para os compostos abaixo.



3) Preveja e explique como a espectroscopia no UV-VIS pode ser utilizada para distinguir os seguintes pares de compostos. Reforce suas respostas com cálculos. Utilize as tabelas fornecidas ao final da Lista.



c)
$$\begin{array}{r} 215 \\ - 12 \\ \hline 227 \text{ nm} \end{array}$$

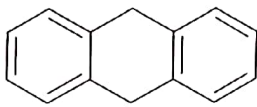


d)
$$\begin{array}{r} 215 \\ + 30 \\ + 18 \\ + 8 \\ \hline 299 \text{ nm} \end{array}$$

a)
$$\begin{array}{r} 215 \\ + 2 \times 12\beta \quad 24 \\ + 1 \times \alpha \quad 5 \\ \hline 244 \text{ nm} \end{array}$$

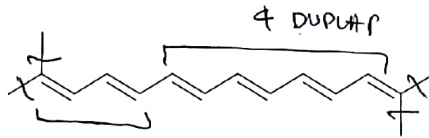
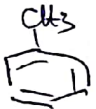
b)
$$\begin{array}{r} 215 \\ - 12 \\ \hline 227 \text{ nm} \end{array}$$

4) Quais dos compostos abaixo deve apresentar absorção na região visível do espectro eletromagnético?

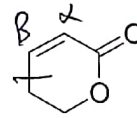


a)
$$\begin{array}{r} 254 \\ - 2 \\ \hline 252 \end{array}$$

$$\begin{array}{r} 261 \\ - 3 \\ \hline 264 \text{ nm} \quad (2 \times \text{max} \\ \text{intensa}) \end{array}$$



b)
$$\begin{array}{r} 217 \\ + 120 \\ + 20 \\ \hline 357 \end{array}$$



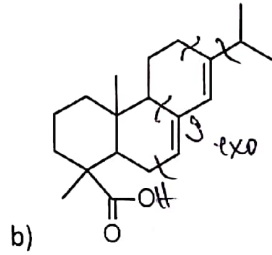
c)
$$\begin{array}{r} 215 \\ - 12 \\ \hline 227 \end{array}$$



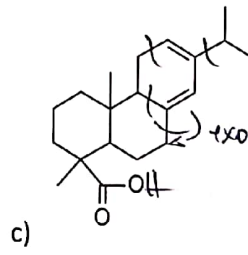
d) Não é planar. As regras de W-F não poderiam ser aplicadas!

5) Calcule os máximos de absorção para os compostos abaixo. Utilize as tabelas fornecidas a seguir.

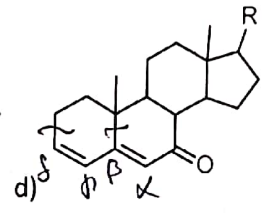
a) 2,4-dimetil-1,3-pentadieno



$$\begin{array}{r}
 214 \\
 4 \times 20 \\
 + 40 \quad 5 \\
 \hline
 239 \text{ nm}
 \end{array}$$



$$\begin{array}{r}
 253 \\
 4 \times 20 \\
 + 40 \quad 5 \\
 \hline
 278 \text{ nm}
 \end{array}$$



$$\begin{array}{r}
 215 \\
 30 \\
 12 \\
 18 \\
 + 40 \quad 5 \\
 \hline
 280 \text{ nm}
 \end{array}$$