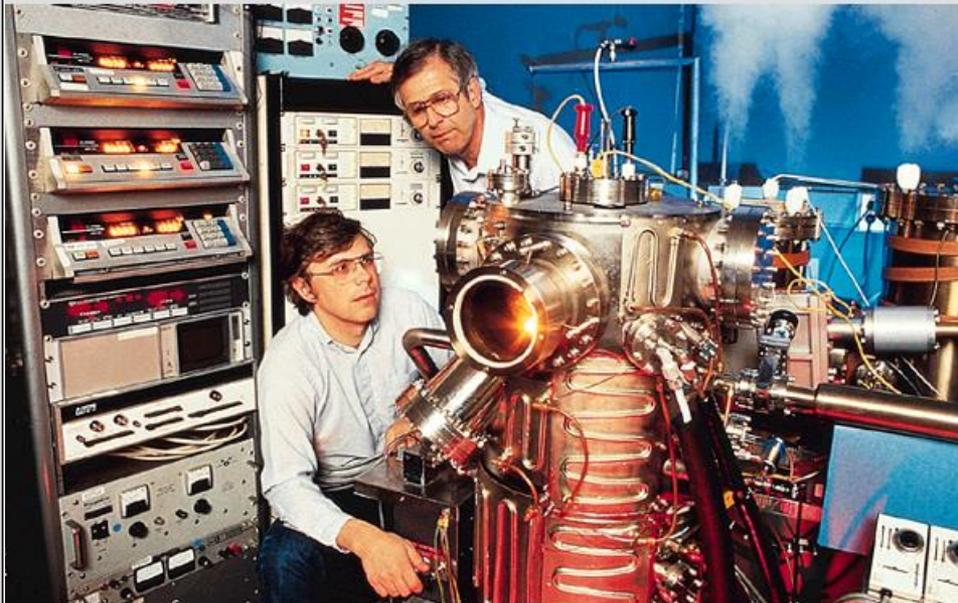


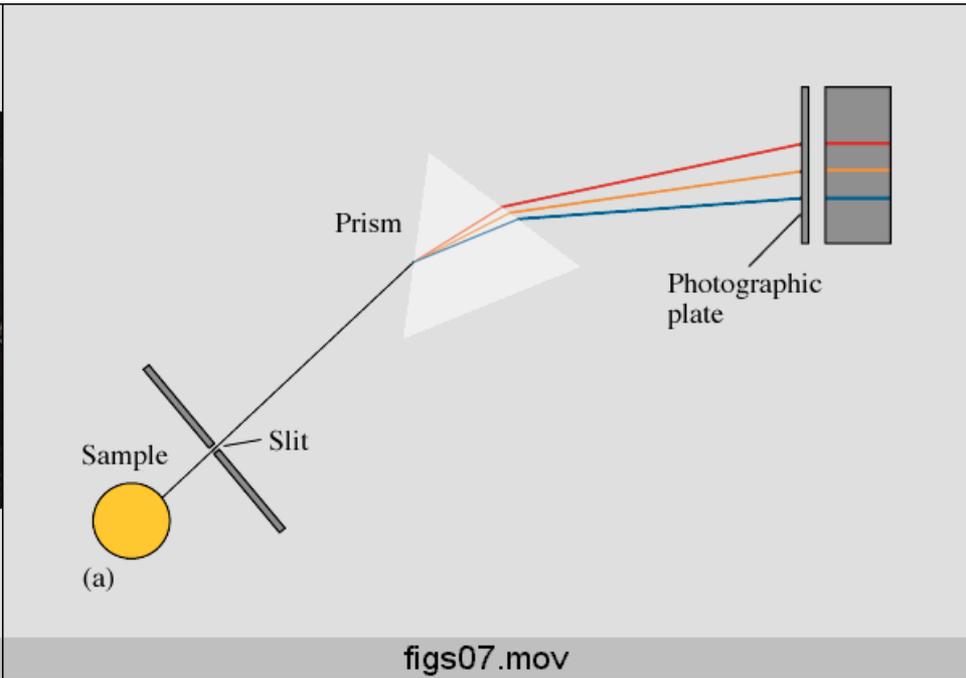
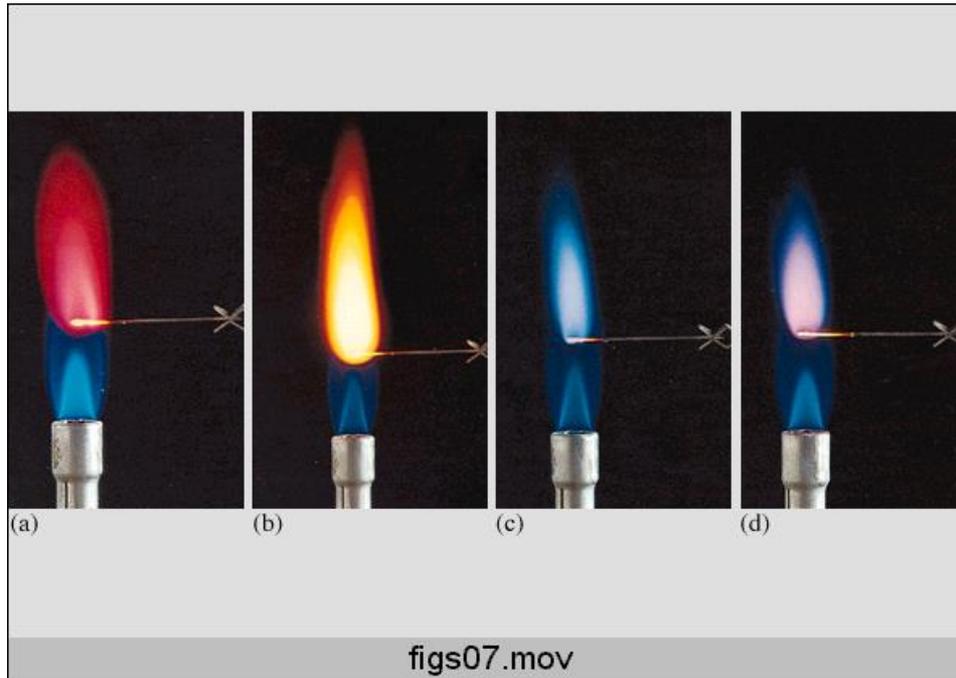
# Estrutura Atômica - Tabela Periódica - Ligações Químicas



figs07.mov



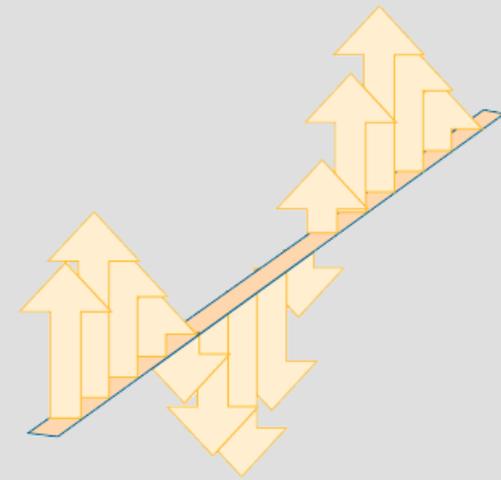
figs07.mov



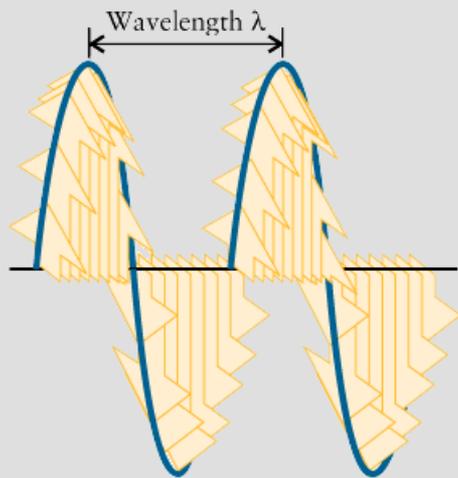


(b)

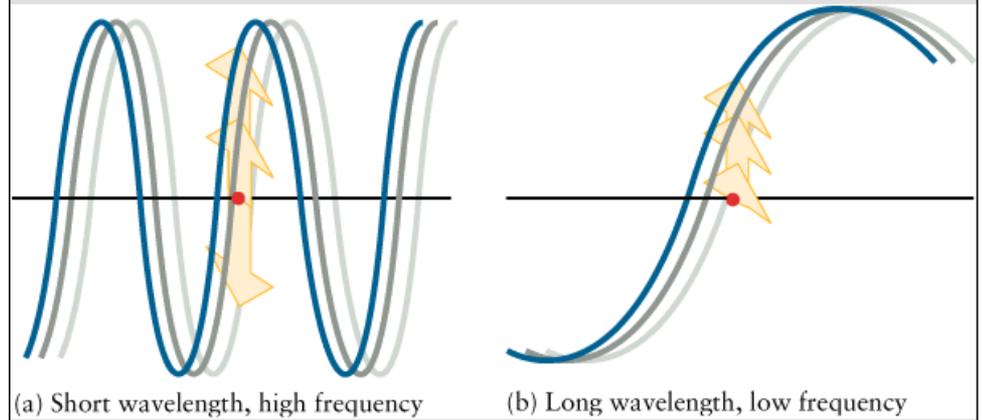
figs07.mov



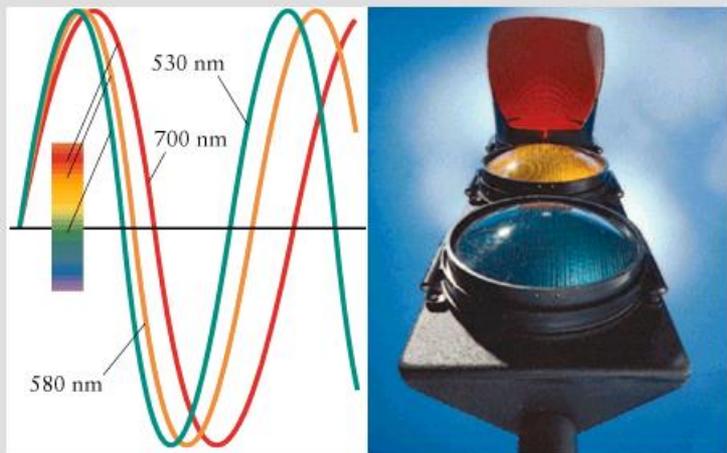
figs07.mov



figs07.mov

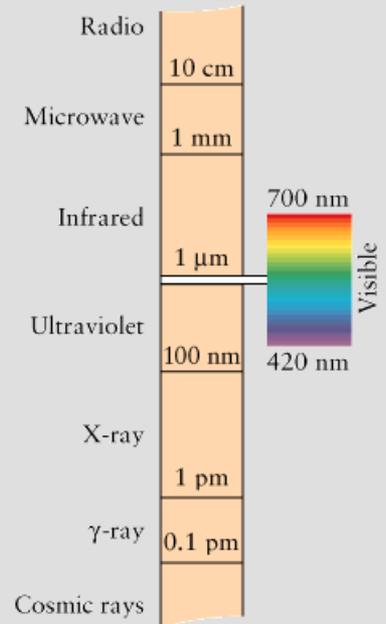


figs07.mov

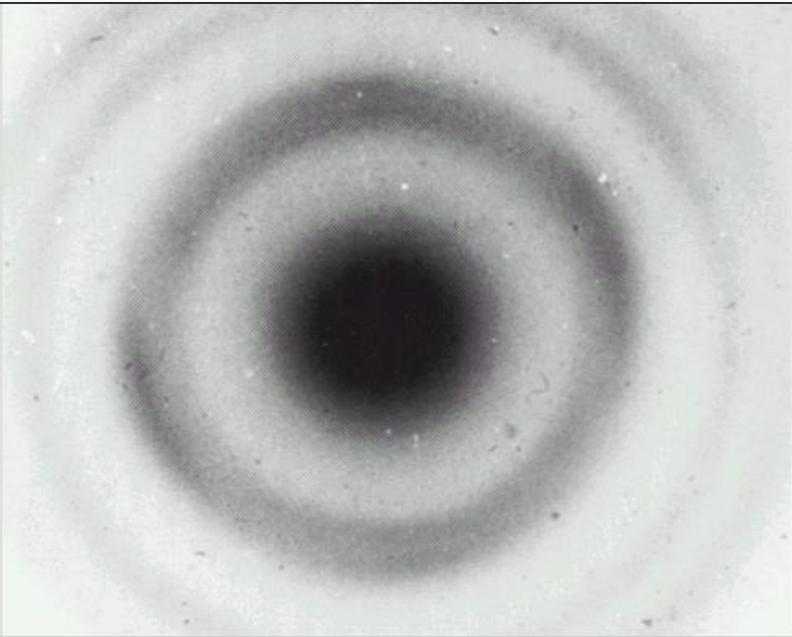


(a) (b)

figs07.mov



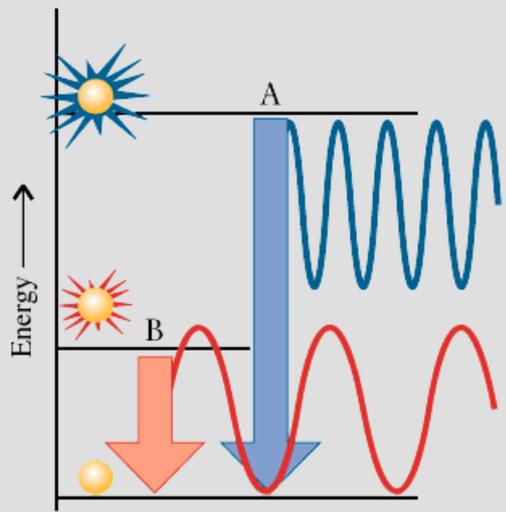
figs07.mov



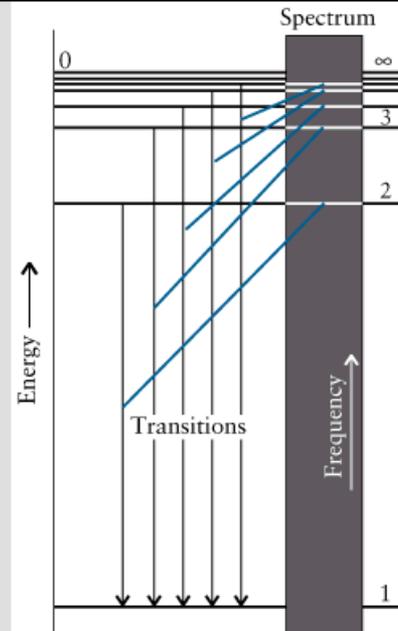
figs07.mov



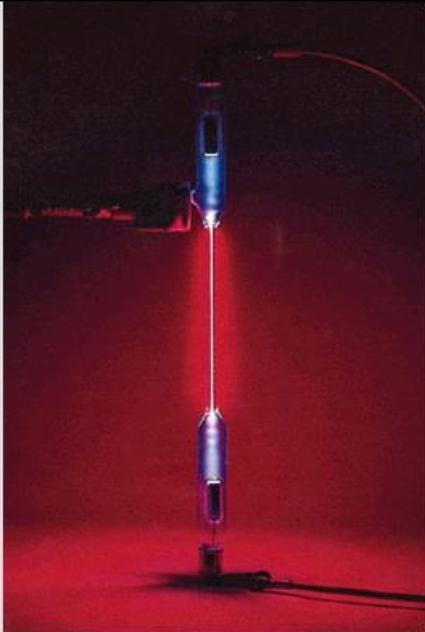
figs07.mov



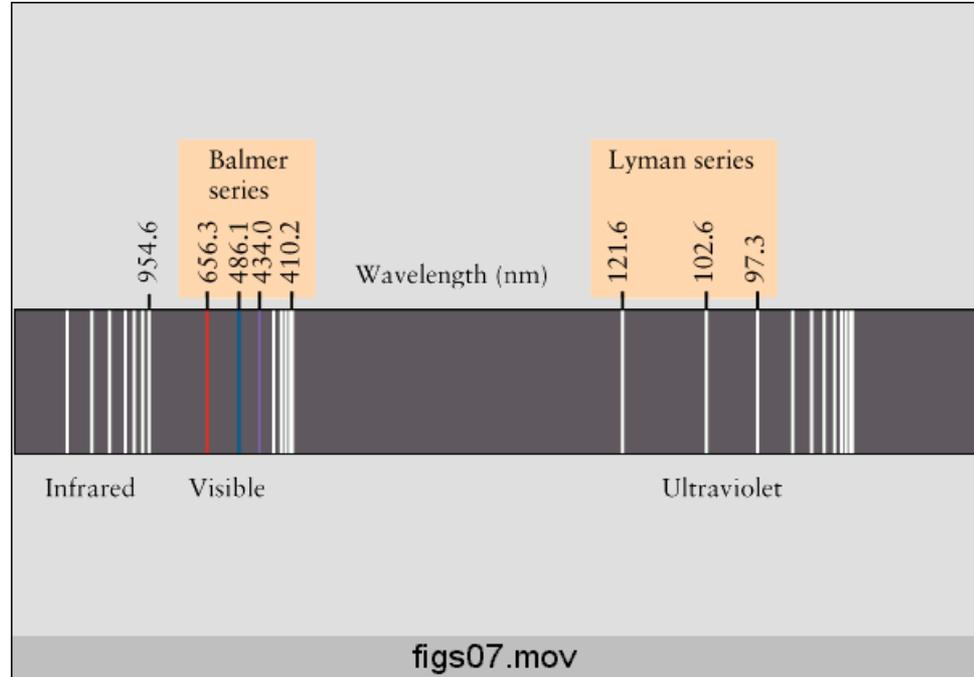
figs07.mov



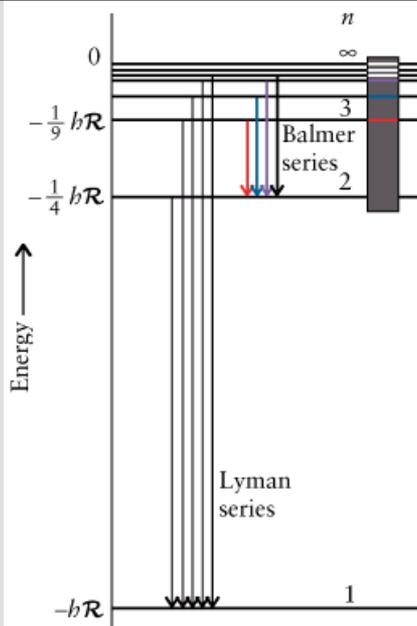
figs07.mov



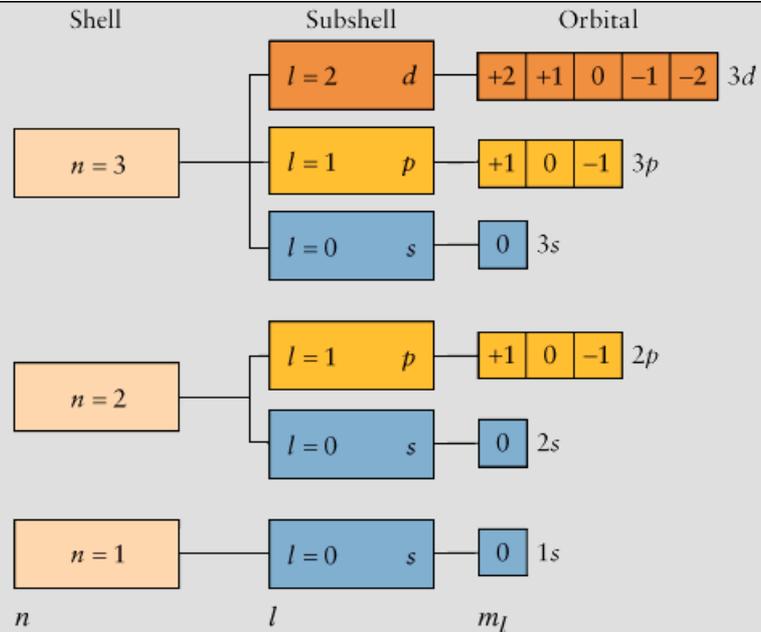
figs07.mov



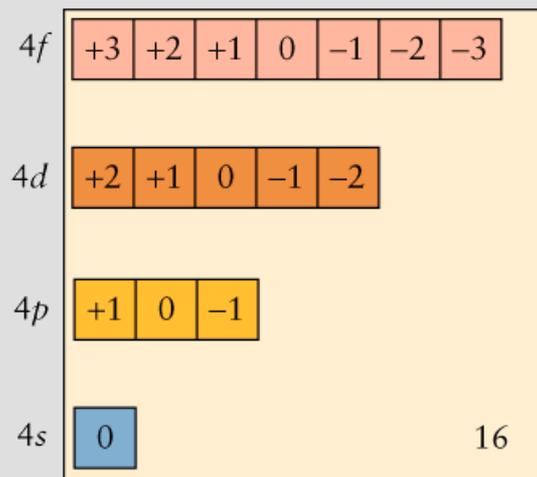
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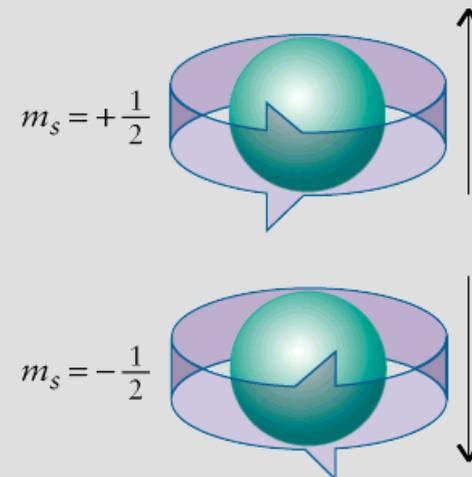
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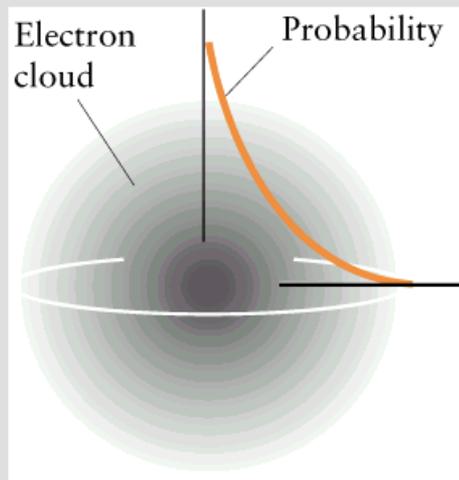
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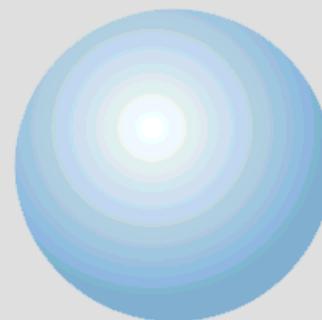
figs07.mov



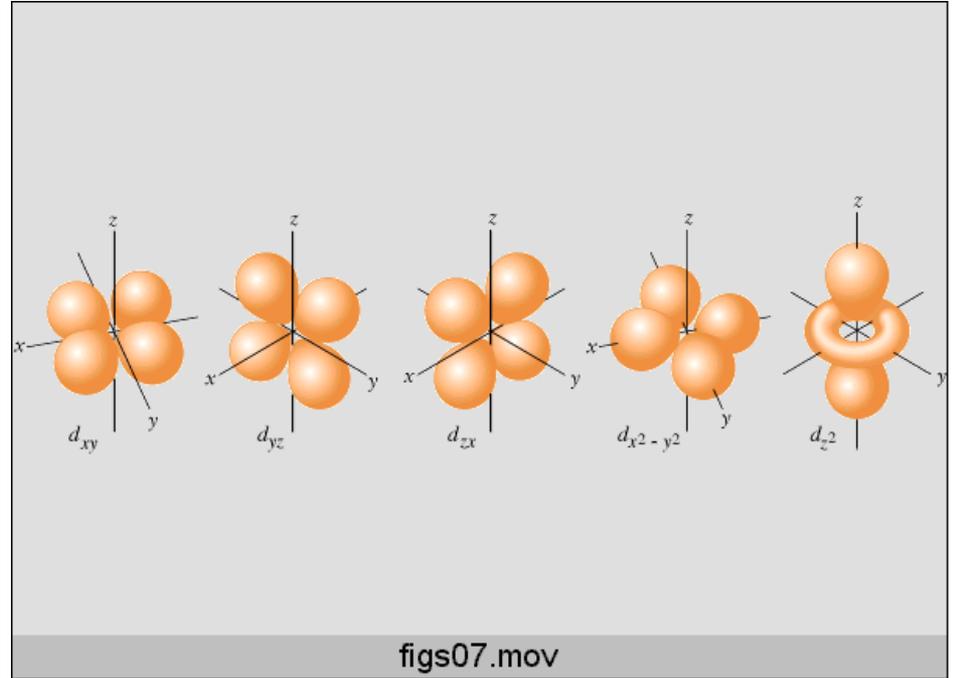
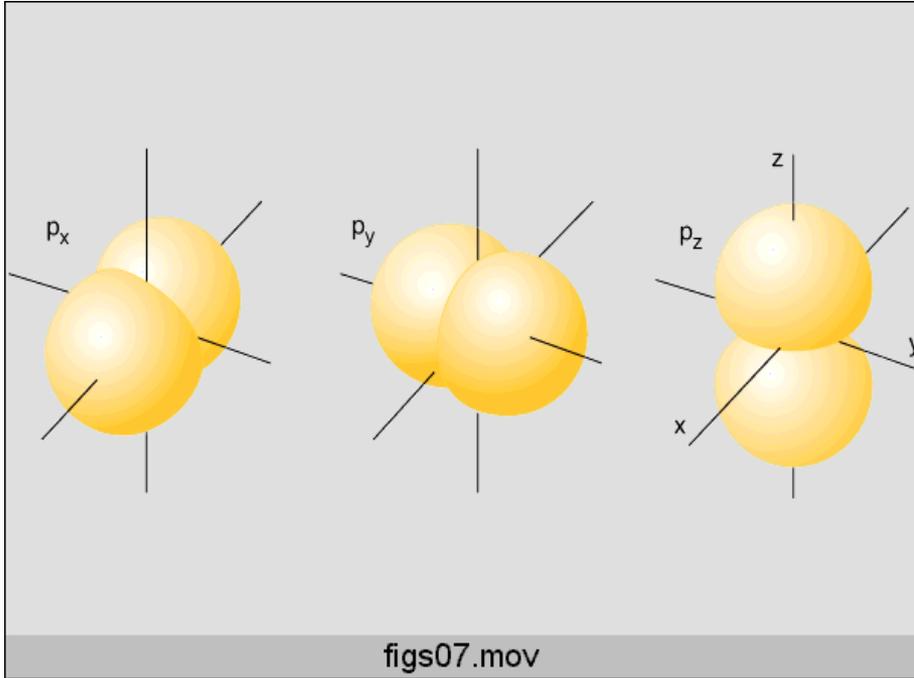
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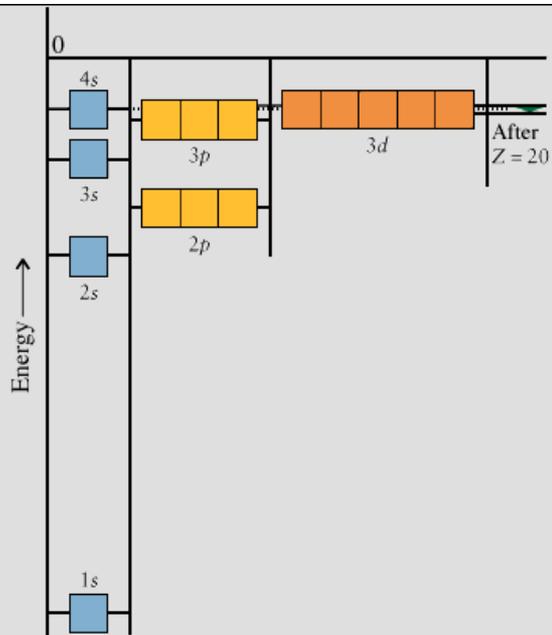


figs07.mov

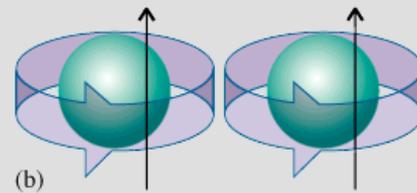
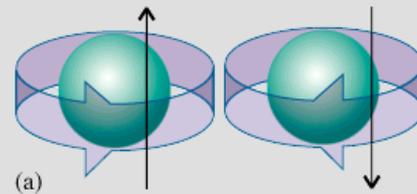


figs07.mov



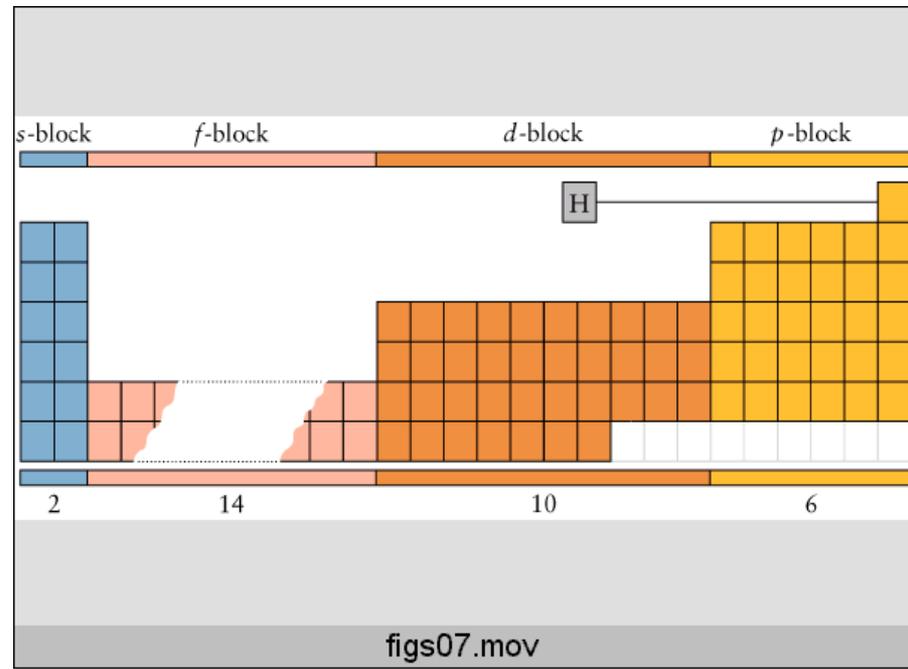
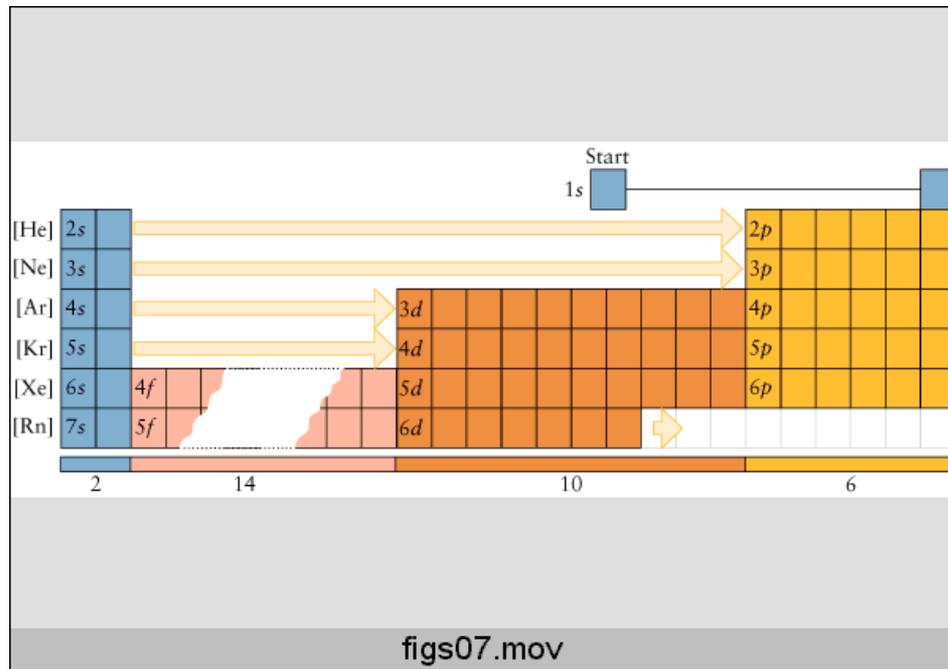


figs07.mov

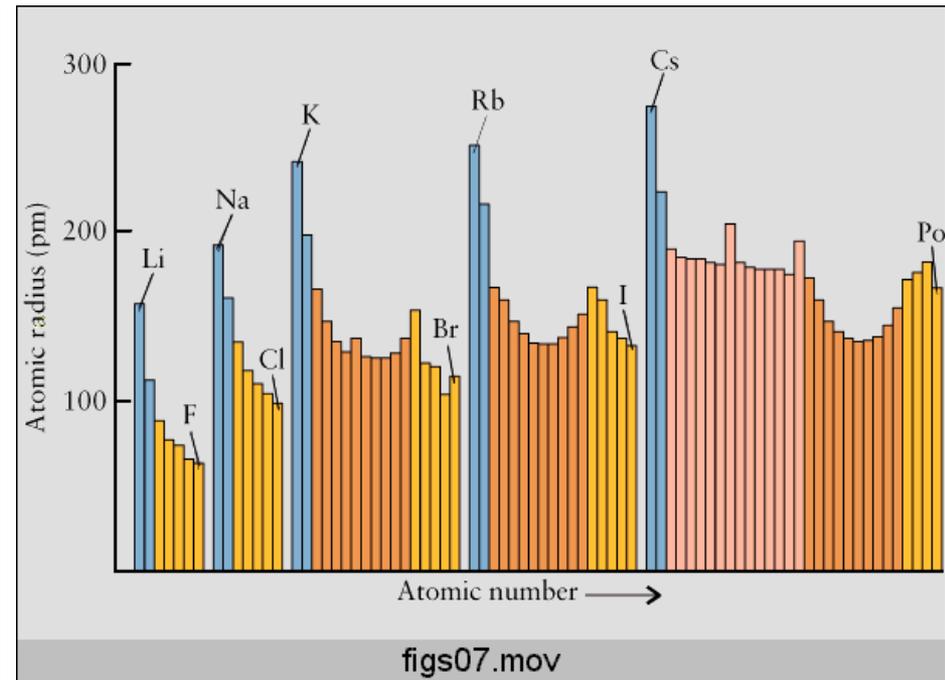
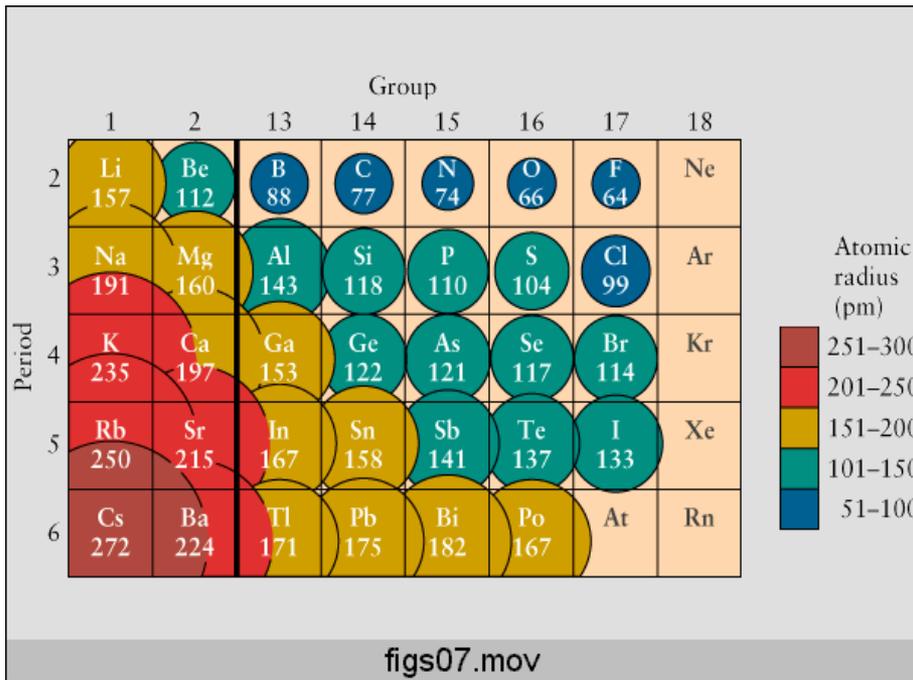
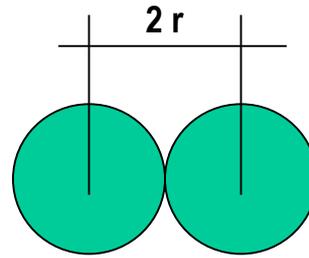


figs07.mov

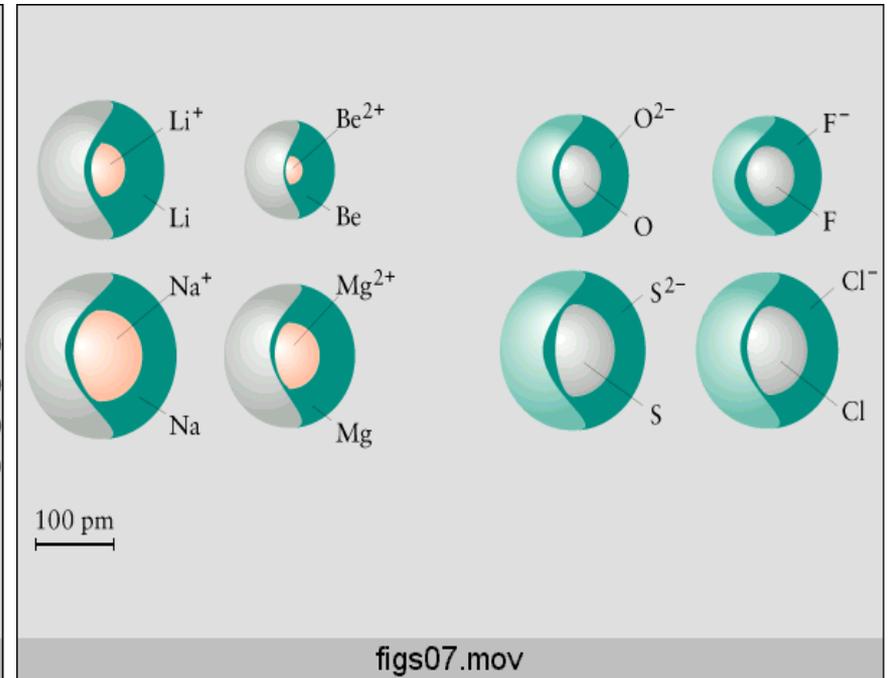
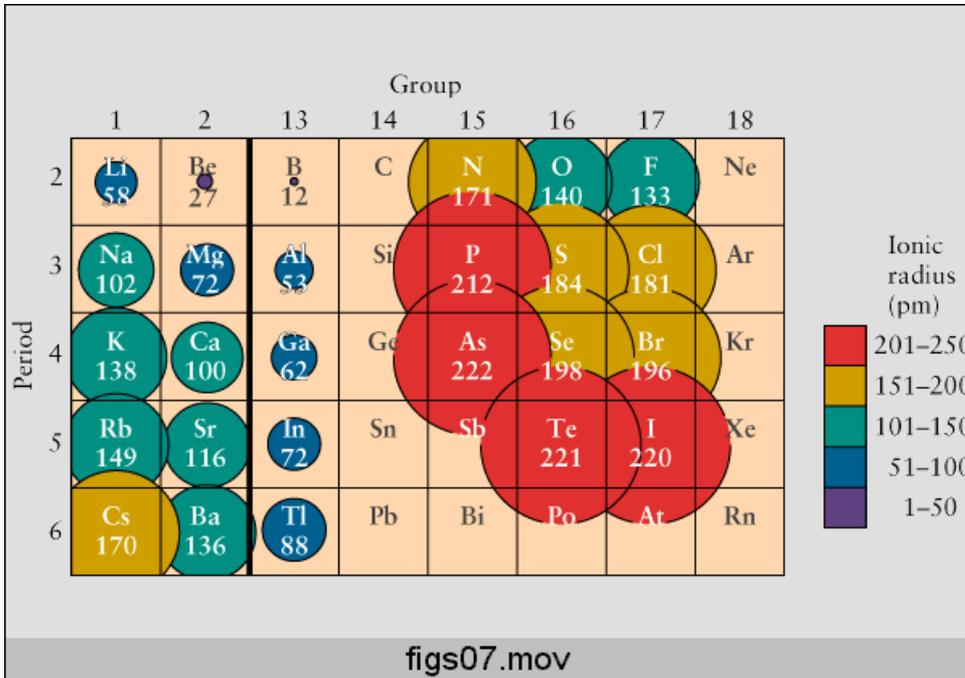
# **A estrutura eletrônica e a Tabela Periódica**



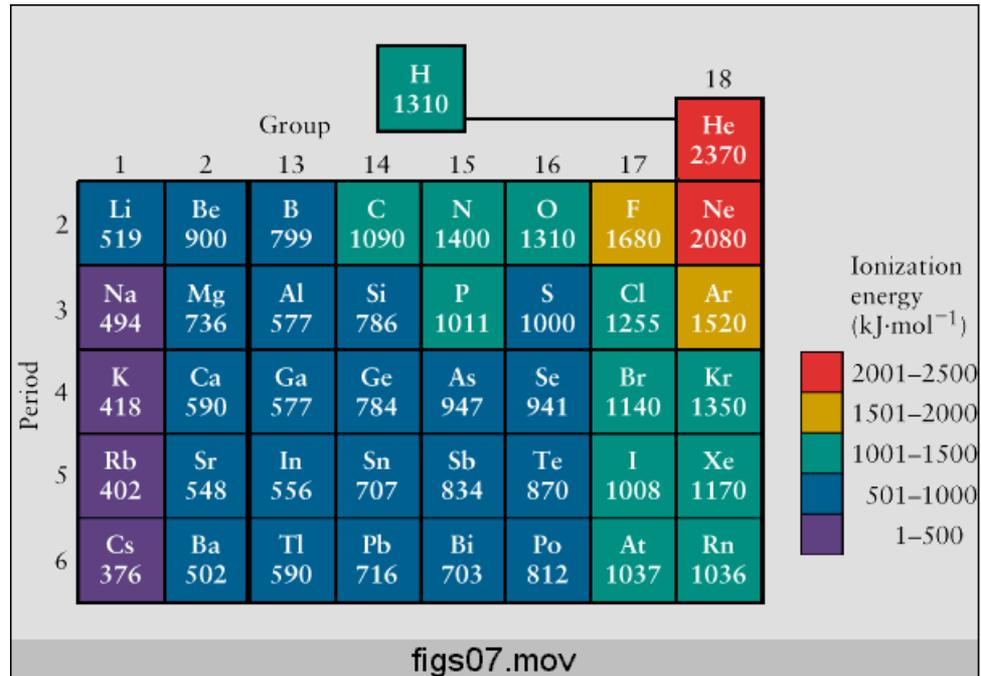
# O raio atômico



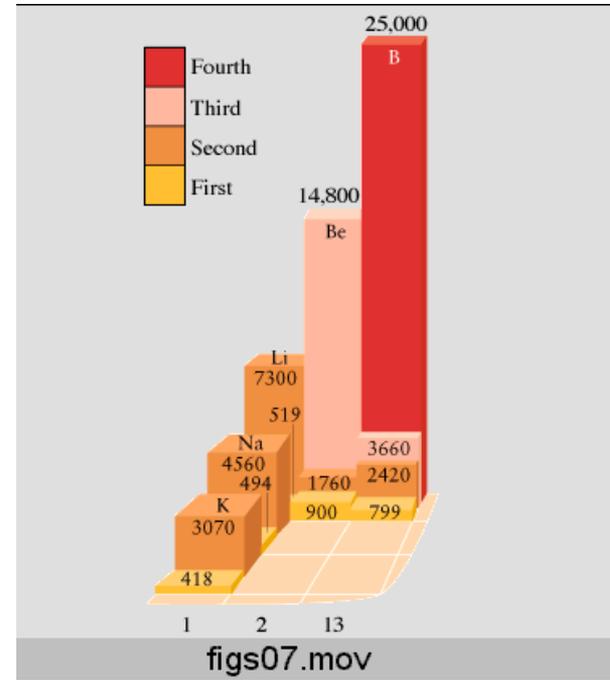
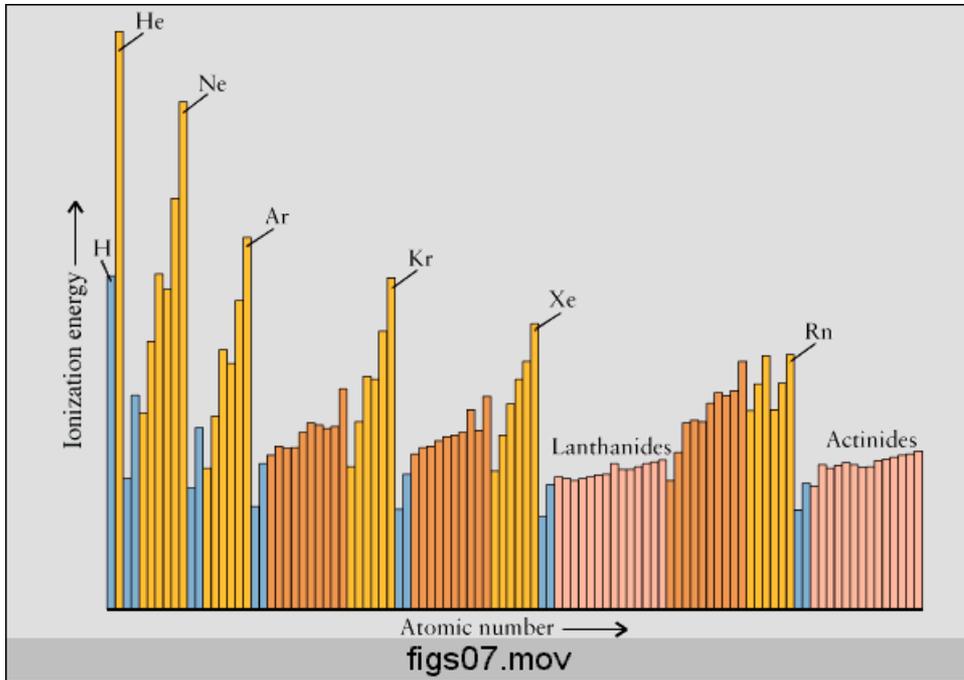
# O raio iônico



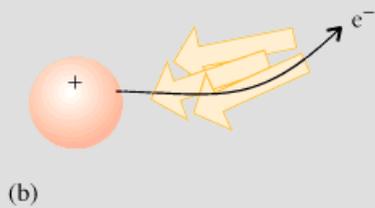
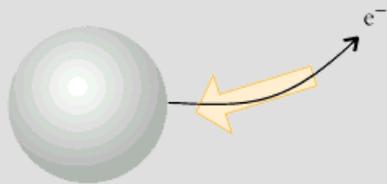
# A energia de ionização



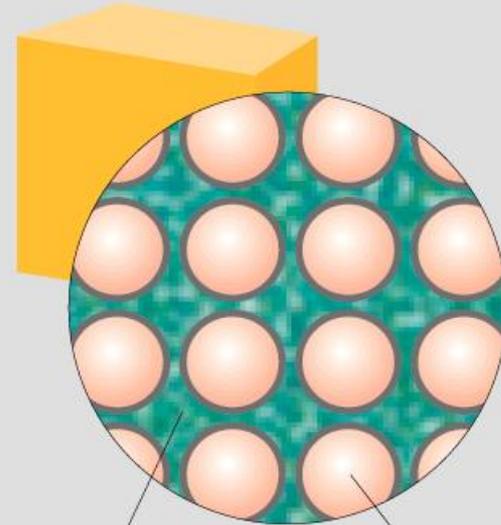
# A energia de ionização



# O efeito do par inerte



figs07.mov



Electron sea

Cation

figs07.mov

# O efeito do par inerte e as relações diagonais



figs07.mov

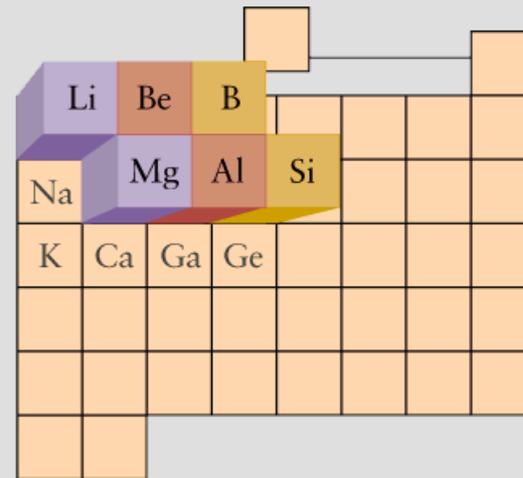
	13	14	15	16
	Al	Si	P	S
Zn	Ga	Ge	As	Se
	In <sup>+</sup> In <sup>3+</sup>	Sn <sup>2+</sup> Sn <sup>4+</sup>	Sb <sup>3+</sup> Sb <sup>5+</sup>	
Cd				
	Tl <sup>+</sup> Tl <sup>3+</sup>	Pb <sup>2+</sup> Pb <sup>4+</sup>	Bi <sup>3+</sup> Bi <sup>5+</sup>	
Hg				

figs07.mov

# O efeito do par inerte e as relações diagonais



figs07.mov



figs07.mov

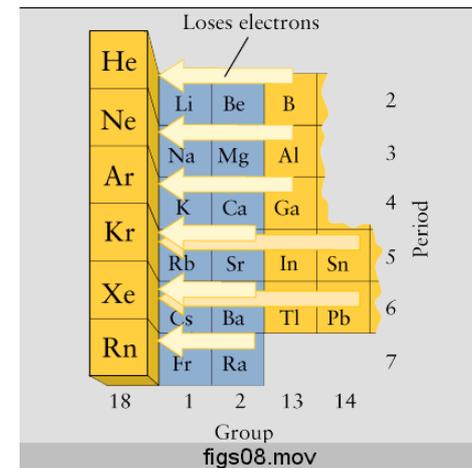
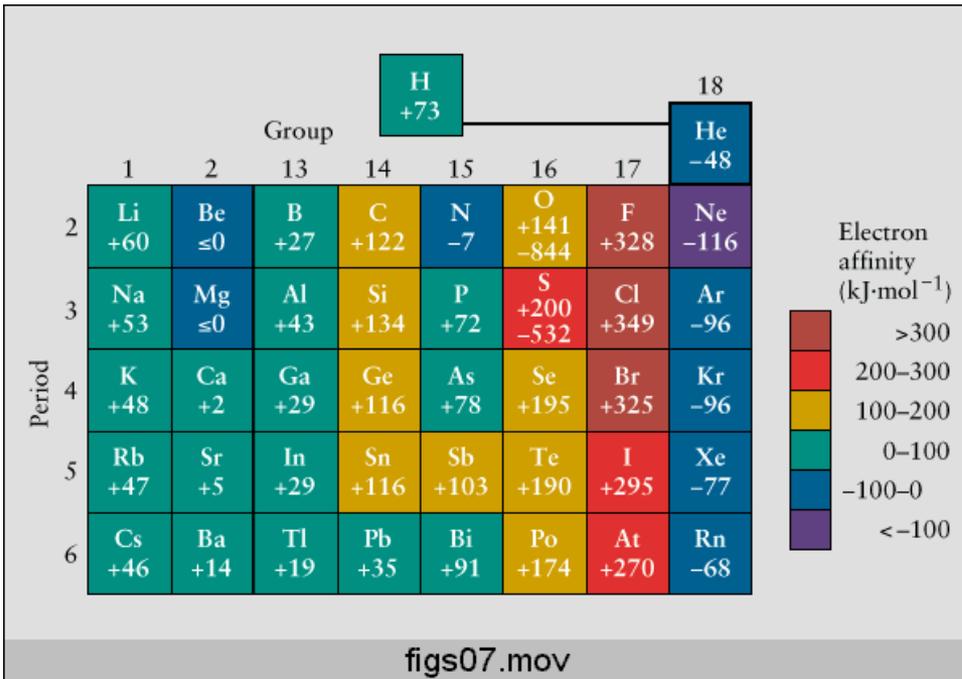
# A afinidade eletrônica

$$E_{ca} = E(X) - E(X^-)$$

Grupo 17 – F, Cl, Br, I (-1)

Grupo 16 – O, S (-2)

# Os elementos dos grupos principais



# Os elementos dos grupos principais

Period	Group							18
	1	2	13	14	15	16	17	
2			B	Si				
3			Al	Ge	As			
4			Ga		Sb	Te		
5			In	Sn		Po		
6				Pb	Bi			
7								

figs07.mov



figs07.mov

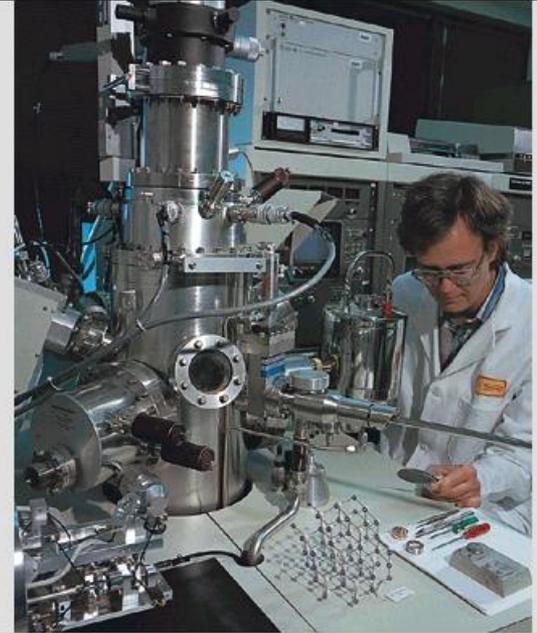
# Os metais de transição



# Ligações Químicas

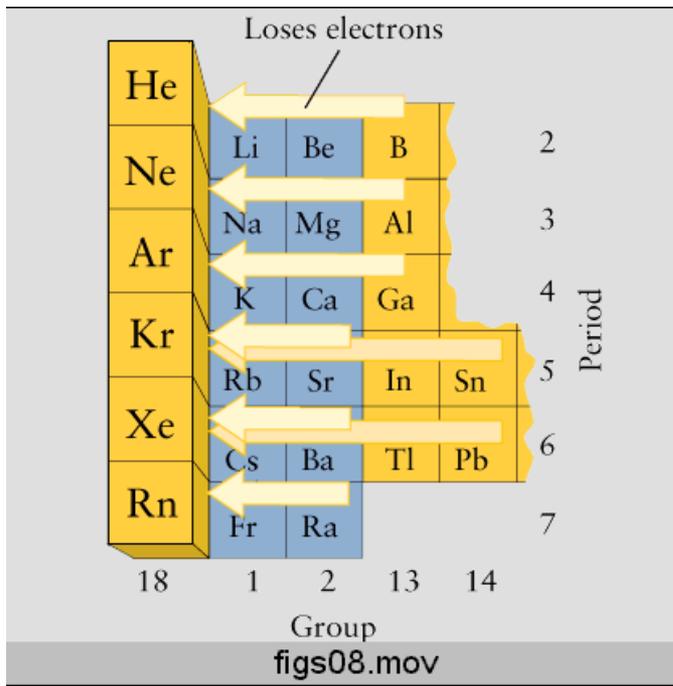
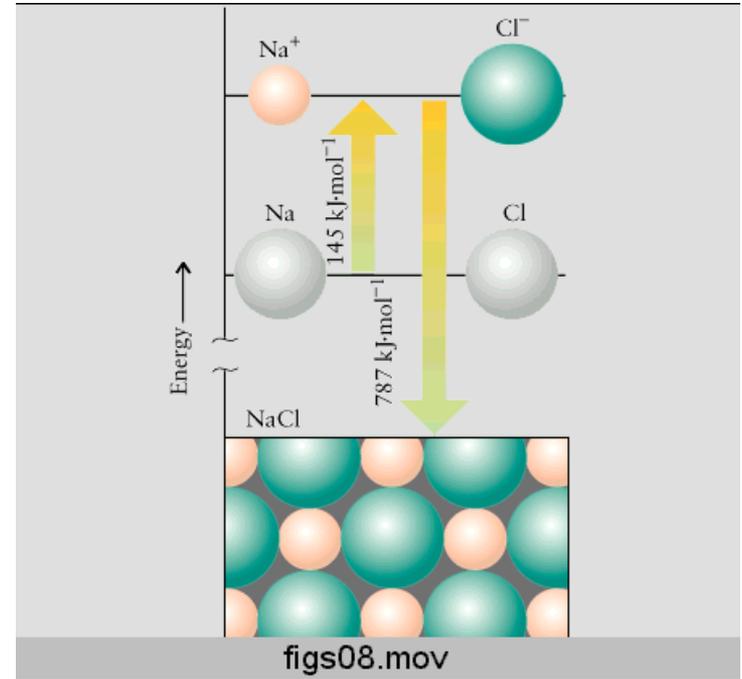
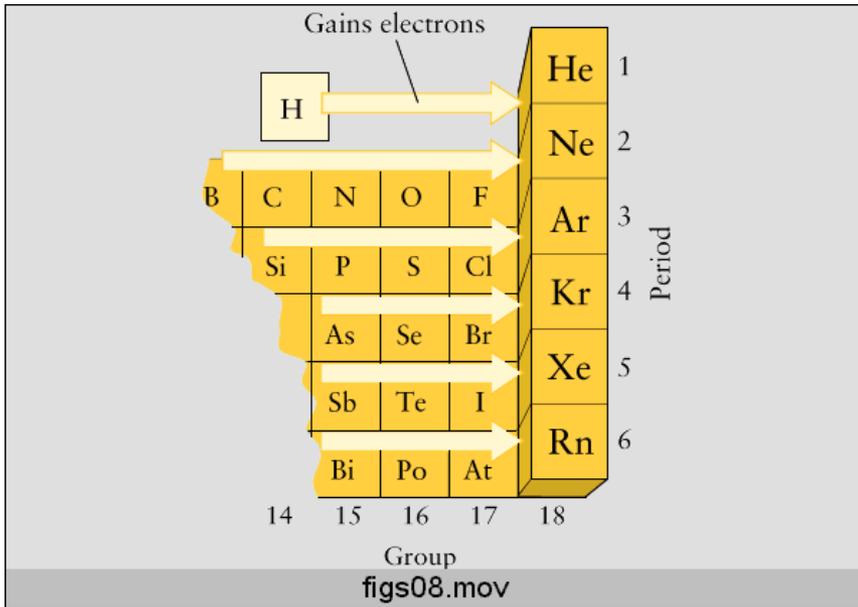


figs08.mov

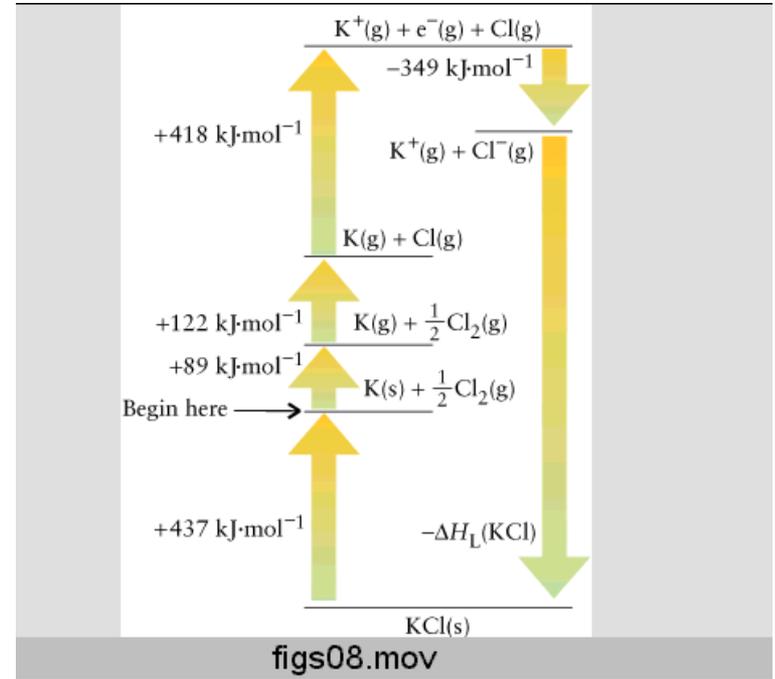
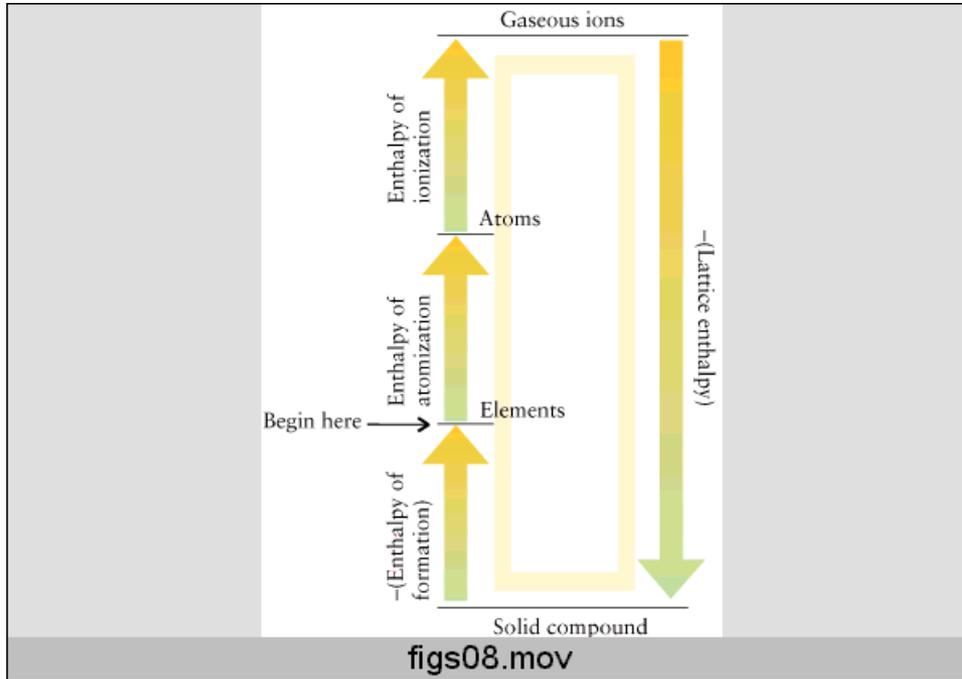


figs08.mov

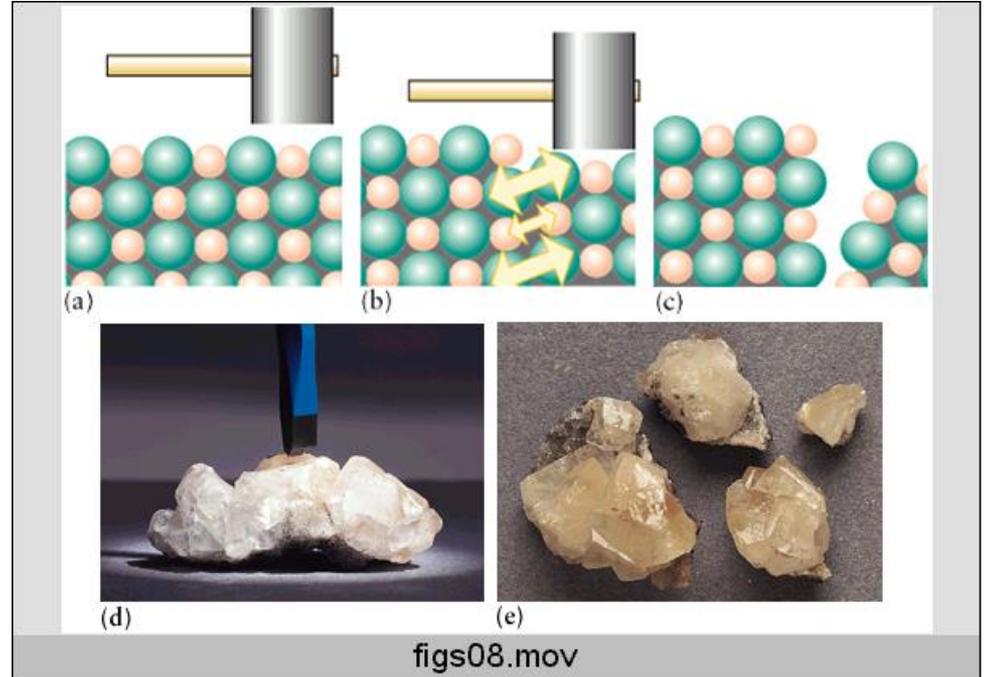
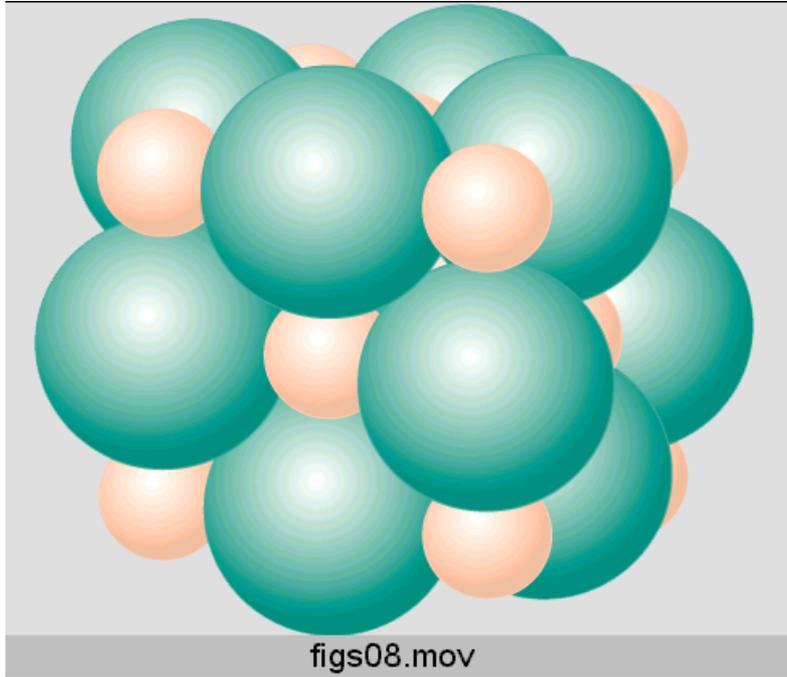
# Ligação iônica



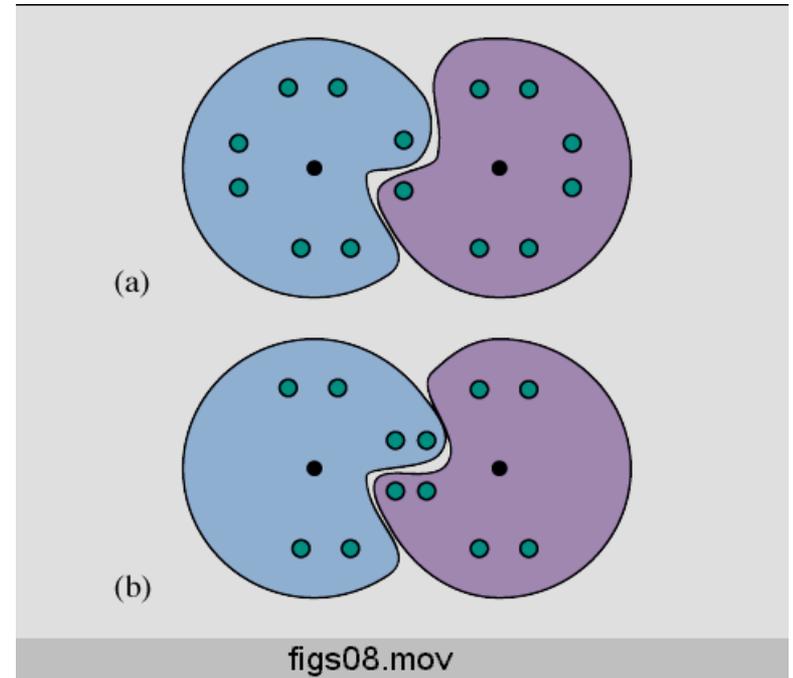
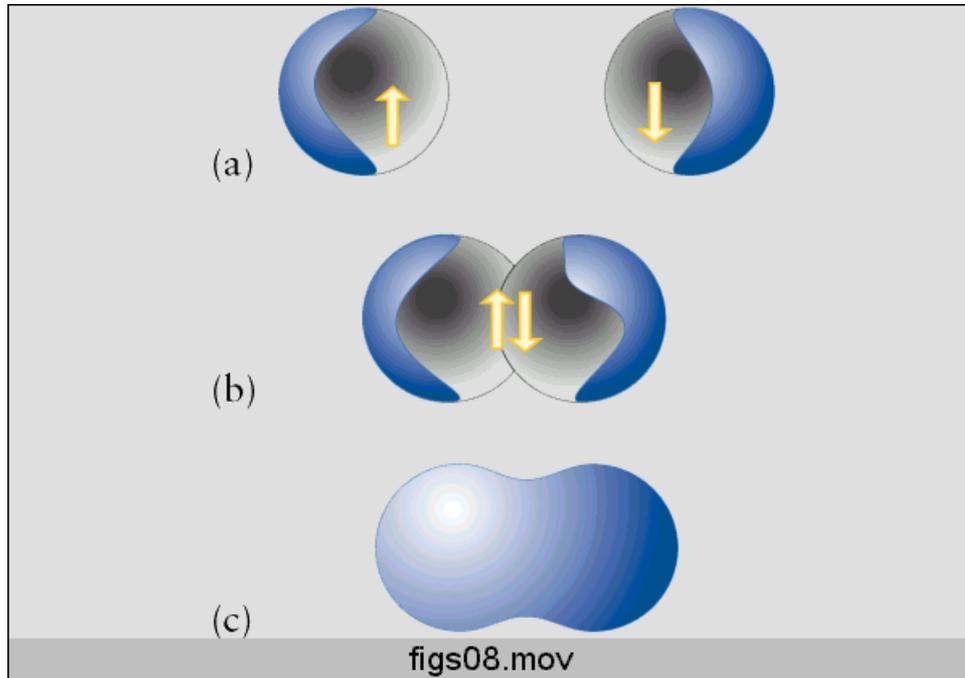
# Ligação iônica



# Ligação iônica



# Ligação covalente



Estrutura de Lewis – Regra do Octeto

# O poder de atração entre os átomos

